

**Chemistry | Semester 2 | Final Exam Study Guide**  
**Unit 9: Compound Stoichiometry**

38. A mole is a unit defined to be  $= 6.02 \times 10^{23}$  particles  
 1 mole = mass

39. The molar mass of an element is the mass of one mole of the substance  
 = its atomic mass in grams

40. What information is given to you in the empirical formula?

- elements in comp.
- ratio of elements

41. What do the subscripts represent in a molecular formula?  
 how many atoms of each element

42. Define Formula Mass  
 mass of 1 mole of an ionic compound

43. Define Molar Mass  
 mass of 1 mole of any compound

44. Define Molecular Mass  
 mass of one mole of a covalent (molecular) comp.

True/False

45. F One mole of  $O_2$  has a greater mass than 1 mole of  $Na_2O$ .  
 32 vs 61.98

46. F One mole of  $H_2$  has <sup>same</sup> molecules than 1 mole of  $Na_2O$ .  
 same # molecules

47. F One molecule of  $H_2$  has more atoms than 1 molecule of  $Na_2O$ .  
 2 atoms vs 3 atoms

**Calculations**

48. Find the molar mass of  $Al_2(SO_4)_3$

Find mass of each element  $\rightarrow$  multiply each mass by subscript  $\rightarrow$  add masses of all elements together = molar mass

Al (2)  $26.98 = 53.96$   
 S (3)  $32.07 = 96.21$   
 O (12)  $16.00 = 192.00$   
 294.17 g/mol

49. How many moles are in 2.4g of carbon dioxide ( $CO_2$ )?

Start with the given value  $\rightarrow$  use dimensional analysis so you can factor out labels until you get to target units

$2.4g CO_2 \times \frac{1 \text{ mol } CO_2}{44 \text{ g } CO_2} = 0.05 \text{ mol}$   
 molar mass  $CO_2$

50. How many molecules are contained in 35.37g of water (molar mass of water is 18.02g/mol)?

$35.37g H_2O \times \frac{1 \text{ mol } H_2O}{18.02g H_2O} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol } H_2O} = 1.18 \times 10^{24} \text{ molecules } H_2O$

51. How many atoms are in 6.3 moles of calcium?

Hint: there are  $6.02 \times 10^{23}$  atoms in 1 mole

$6.3 \text{ mol Ca} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol Ca}} = 3.79 \times 10^{24} \text{ atoms}$

52. Calculate the number of moles there are in  $4.5 \times 10^{28}$  molecules of sulfur dioxide ( $SO_2$ ).

$4.5 \times 10^{28} \text{ molecules } SO_2 \times \frac{1 \text{ mol } SO_2}{6.02 \times 10^{23} \text{ molecules}} = 747.51 \text{ mol } SO_2$

53. Calculate the percent composition of sodium and oxygen in sodium oxide ( $Na_2O$ ).

Find the molar mass  $\rightarrow$  divide total mass of element where it's solving for by the total mass of compound

Na (2)  $22.99 = 45.98 / 61.98 = 74.2\%$   
 O (1)  $16.00 = 16.00 / 61.98 = 25.8\%$

molar mass

45.98  
 16.00  
 61.98 g/mol

54. What is the empirical formula of a compound that contains 81.68% carbon and 18.32% hydrogen?

Assume 100 g: % to mass  $\rightarrow$  mass to mole  $\rightarrow$  divide by smallest  $\rightarrow$  multiply till whole

$81.68g C \times \frac{1 \text{ mol}}{12.01g C} = 6.8 / 6.8 = 1 \times 3 = 3$   
 $18.32g H \times \frac{1 \text{ mol}}{1.01g H} = 18.139 / 6.8 = 2.67 \times 3 = 8$   
 $C_3H_8$

55. What is the molecular formula of a hydrocarbon that has an empirical formula of  $CH_3$  and a molecular mass (molar mass of molecule) of 48.15 g/mol?

$\frac{MM_{\text{molecule}}}{MM_{\text{empirical}}} = \frac{48.15}{16.05} = 3$   
 $C_3H_{12}$

Multiply each subscript by the above answer to get the subscripts for the molecular formula

To find Empirical Formula  
 divide (multiply till whole)

if it ends in multiply by  
 0.2 or 0.8  $\times 5$   
 0.25 or 0.75  $\times 4$   
 0.33 or 0.67  $\times 3$   
 0.5  $\times 2$

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