

# AP Chemistry Final Study Guide

## Unit 1: Atomic Structure and Properties (7%–9%)

- At STP, one mole of gas always occupies 22.4 L
- **Mass Spectroscopy:**
  - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
  - The most abundant isotope will have the greatest effect
- **Coulomb's Law:**
  - The greater the charge of the nucleus, the more energy is required to ionize an electron
  - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- **Electromagnetic Radiation:**
  - When atoms absorb energy, electrons jump to higher energy levels
  - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- **Kinetic Energy of Ionized Electrons:**
  - The faster an ejected electron is going, the more kinetic energy it has
  - Electrons originally further away from the nucleus require less energy to eject, so it will have more kinetic energy
- **Electron Configuration:**
  - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
    - Nickel:  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^8 4s^2$  (notice the d-block)
    - Nickel (with shorthand):  $[\text{Ar}] 4s^2 3d^8$
- **Law of Definite Proportions:**
  - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- **Law of Multiple Proportions:**
  - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- **Aufbau Principle:**
  - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

# Chemistry Final Exam Study Guide Calculations

**JE Gale**



## **Chemistry Final Exam Study Guide Calculations:**

## Whispering the Techniques of Language: An Emotional Journey through **Chemistry Final Exam Study Guide Calculations**

In a digitally-driven earth where displays reign great and instant connection drowns out the subtleties of language, the profound strategies and psychological nuances concealed within phrases usually get unheard. However, situated within the pages of **Chemistry Final Exam Study Guide Calculations** a captivating fictional value blinking with organic thoughts, lies an extraordinary quest waiting to be undertaken. Penned by an experienced wordsmith, that wonderful opus invites visitors on an introspective journey, lightly unraveling the veiled truths and profound impact resonating within the material of each word. Within the emotional depths of this emotional review, we will embark upon a genuine exploration of the book is key styles, dissect its captivating writing fashion, and succumb to the strong resonance it evokes deep within the recesses of readers hearts.

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