

Answer Key

PG-1: What is the coefficient of the hydroxide ion for the half reaction when the hypochlorite ion is reduced? (Note the reaction is balanced in acidic solution.) unbalanced



PG-2: What is the balanced half reaction of the species in reacting to form nitrogen monoxide in acidic solution?



PG-3: What is the coefficient for water when the reaction is balanced in acidic solution? unbalanced



PG-4: Which reaction occurs in the cell in the presence of standard state conditions?

The standard reduction potential of zinc is  $+1.26 \text{ V}$ .  
The standard reduction potential of iron is  $+0.77 \text{ V}$ .



The cell is to do work, releasing energy. The reduction half reaction is  $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$  with  $E^\circ = +0.77 \text{ V}$ . The oxidation half reaction is  $\text{Zn} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$  with  $E^\circ = +1.26 \text{ V}$ . For a reaction to be spontaneous, the cell potential must be positive. The cell potential is  $E^\circ_{\text{cell}} = E^\circ_{\text{cathode}} - E^\circ_{\text{anode}} = 0.77 \text{ V} - 1.26 \text{ V} = -0.49 \text{ V}$ . Since the cell potential is negative, the reaction is non-spontaneous. The reaction is  $\text{Zn}^{2+} + \text{Fe} \rightarrow \text{Zn} + \text{Fe}^{2+}$ .

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