

Questions 17–19

- (A)  $\text{NaC}_2\text{H}_3\text{O}_2$
- (B)  $\text{HC}_2\text{H}_3\text{O}_2$
- (C)  $\text{KCl}$
- (D)  $\text{NH}_3$
- (E)  $\text{HCl}$

17. Is a salt that will undergo hydrolysis to form a basic solution.
18. Will form a coordinate covalent bond with a hydronium ion.
19. Is a strong acid.

Questions 20–22

- (A)  $q = mc\Delta T$
- (B)  $q = H_m$
- (C)  $P_1V_1 = P_2V_2$
- (D)  $D = m / V$
- (E)  $K = C + 273$

20. Can be used to find the mass of an irregularly shaped solid.
21. Boyle's Law.
22. Used to find energy gained or lost during a particular phase change.

Questions 23–25

- (A) Alpha particle
- (B) Beta particle
- (C) Neutron
- (D) Gamma ray
- (E) Positron

23. Has the greatest mass.
24. Has the greatest positive charge.
25. Has the same mass and charge as an electron.

# Chemistry Diagnostic Test Topic 7 Answers

**Joseph A. Mascetta, Mark Kernion**



## Chemistry Diagnostic Test Topic 7 Answers:

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