

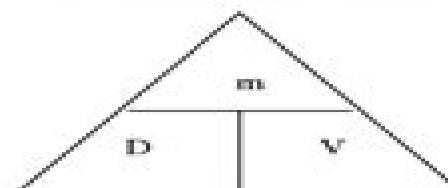
**SNC 1D0**  
**Unit: Chemistry**  
**Density Calculations Worksheet**

Name: \_\_\_\_\_

Date: \_\_\_\_\_

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

UNITS OF DENSITY  
 $\text{g/cm}^3$  or  $\text{g/mL}$



1. Find the unknown quantity.

a.  $D = 3 \text{ g/mL}$   
 $V = 100 \text{ mL}$   
 $m = ?$

300 g

b.  $D = ?$   
 $V = 950 \text{ mL}$   
 $m = 95 \text{ g}$

0.10 g/mL

c.  $D = 0.5 \text{ g/cm}^3$   
 $V = ?$   
 $m = 20 \text{ g}$

40  $\text{cm}^3$

2. Find the unknown quantity (CONVERT FIRST to g or mL).

a.  $D = 24 \text{ g/mL}$   
 $V = 1.2 \text{ L} = \underline{\hspace{1cm}} 1200 \underline{\hspace{1cm}} \text{ mL}$   
 $M = ?$

28800 g = 28.8 kg

b.  $D = ?$   
 $V = 100 \text{ mL}$   
 $M = 1.5 \text{ kg} = \underline{\hspace{1cm}} 1500 \underline{\hspace{1cm}} \text{ g}$

15 g/mL

c.  $D = ?$   
 $V = 0.52 \text{ L} = \underline{\hspace{1cm}} 520 \underline{\hspace{1cm}} \text{ mL}$   
 $M = 500 \text{ mg} = \underline{\hspace{1cm}} 0.5 \underline{\hspace{1cm}} \text{ g}$

0.00096 g/mL

**WORD PROBLEMS**

1. A block of aluminum occupies a volume of 15.0 mL and has a mass of 40.5 g. What is its density?

2.70 g/mL

2. Mercury metal is poured into a graduated cylinder that holds exactly 22.5 mL. The mercury is used to fill the cylinder has a mass of 306.0 g. From this information, calculate the density of mercury.

# Chemistry Density Word Problems Worksheet

**Jin-Ying Zhang**



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