

HIGHER SCHOOL CERTIFICATE EXAMINATION Chemistry

DATA SHEET

Avogadro constant, N_A	$6.022 \times 10^{23} \text{ mol}^{-1}$
Volume of 1 mole ideal gas: at 100 kPa and	
at 0°C (273.15 K)	22.71 L
at 25°C (298.15 K)	24.79 L
Ionisation constant for water at 25°C (298.15 K), K_w	1.0×10^{-14}
Specific heat capacity of water	$4.18 \times 10^3 \text{ J kg}^{-1} \text{ K}^{-1}$

Some useful formulae

$$\text{pH} = -\log_{10}[\text{H}^+]$$

$$\Delta H = -m C \Delta T$$

Some standard potentials

$\text{K}^+ + \text{e}^-$	e^0	$\text{K}(\text{s})$	-2.94 V
$\text{Ba}^{2+} + 2\text{e}^-$	e^0	$\text{Ba}(\text{s})$	-2.91 V
$\text{Ca}^{2+} + 2\text{e}^-$	e^0	$\text{Ca}(\text{s})$	-2.87 V
$\text{Na}^+ + \text{e}^-$	e^0	$\text{Na}(\text{s})$	-2.71 V
$\text{Mg}^{2+} + 2\text{e}^-$	e^0	$\text{Mg}(\text{s})$	-2.36 V
$\text{Al}^{3+} + 3\text{e}^-$	e^0	$\text{Al}(\text{s})$	-1.68 V
$\text{Mn}^{2+} + 2\text{e}^-$	e^0	$\text{Mn}(\text{s})$	-1.18 V
$\text{H}_2\text{O} + \text{e}^-$	e^0	$\frac{1}{2}\text{H}_2(\text{g}) + \text{OH}^-$	-0.83 V
$\text{Zn}^{2+} + 2\text{e}^-$	e^0	$\text{Zn}(\text{s})$	-0.76 V
$\text{Fe}^{2+} + 2\text{e}^-$	e^0	$\text{Fe}(\text{s})$	-0.44 V
$\text{Ni}^{2+} + 2\text{e}^-$	e^0	$\text{Ni}(\text{s})$	-0.24 V
$\text{Sn}^{2+} + 2\text{e}^-$	e^0	$\text{Sn}(\text{s})$	-0.14 V
$\text{Pb}^{2+} + 2\text{e}^-$	e^0	$\text{Pb}(\text{s})$	-0.13 V
$\text{H}^+ + \text{e}^-$	e^0	$\frac{1}{2}\text{H}_2(\text{g})$	0.00 V
$\text{SO}_4^{2-} + 4\text{H}^+ + 2\text{e}^-$	e^0	$\text{SO}_2(\text{aq}) + 2\text{H}_2\text{O}$	0.16 V
$\text{Cu}^{2+} + 2\text{e}^-$	e^0	$\text{Cu}(\text{s})$	0.34 V
$\frac{1}{2}\text{O}_2(\text{g}) + \text{H}_2\text{O} + 2\text{e}^-$	e^0	2OH^-	0.40 V
$\text{Cu}^+ + \text{e}^-$	e^0	$\text{Cu}(\text{s})$	0.52 V
$\frac{1}{2}\text{I}_2(\text{s}) + \text{e}^-$	e^0	I^-	0.54 V
$\frac{1}{2}\text{I}_2(\text{aq}) + \text{e}^-$	e^0	I^-	0.62 V
$\text{Fe}^{3+} + \text{e}^-$	e^0	Fe^{2+}	0.77 V
$\text{Ag}^+ + \text{e}^-$	e^0	$\text{Ag}(\text{s})$	0.80 V
$\frac{1}{2}\text{Br}_2(\text{l}) + \text{e}^-$	e^0	Br^-	1.08 V
$\frac{1}{2}\text{Br}_2(\text{aq}) + \text{e}^-$	e^0	Br^-	1.10 V
$\frac{1}{2}\text{O}_2(\text{g}) + 2\text{H}^+ + 2\text{e}^-$	e^0	H_2O	1.23 V
$\frac{1}{2}\text{Cl}_2(\text{g}) + \text{e}^-$	e^0	Cl^-	1.36 V
$\frac{1}{2}\text{Cr}_2\text{O}_7^{2-} + 7\text{H}^+ + 3\text{e}^-$	e^0	$\text{Cr}^{3+} + \frac{7}{2}\text{H}_2\text{O}$	1.36 V
$\frac{1}{2}\text{Cl}_2(\text{aq}) + \text{e}^-$	e^0	Cl^-	1.40 V
$\text{MnO}_4^- + 8\text{H}^+ + 5\text{e}^-$	e^0	$\text{Mn}^{2+} + 4\text{H}_2\text{O}$	1.51 V
$\frac{1}{2}\text{F}_2(\text{g}) + \text{e}^-$	e^0	F^-	2.89 V

Chemistry Data Sheet For Ch1fp

Joseph H. Noggle



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Table of Contents Chemistry Data Sheet For Ch1fp

1. Understanding the eBook Chemistry Data Sheet For Ch1fp
 - The Rise of Digital Reading Chemistry Data Sheet For Ch1fp
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Data Sheet For Ch1fp
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Data Sheet For Ch1fp
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Data Sheet For Ch1fp
 - Personalized Recommendations
 - Chemistry Data Sheet For Ch1fp User Reviews and Ratings
 - Chemistry Data Sheet For Ch1fp and Bestseller Lists
5. Accessing Chemistry Data Sheet For Ch1fp Free and Paid eBooks
 - Chemistry Data Sheet For Ch1fp Public Domain eBooks
 - Chemistry Data Sheet For Ch1fp eBook Subscription Services
 - Chemistry Data Sheet For Ch1fp Budget-Friendly Options
6. Navigating Chemistry Data Sheet For Ch1fp eBook Formats

- ePub, PDF, MOBI, and More
- Chemistry Data Sheet For Ch1fp Compatibility with Devices
- Chemistry Data Sheet For Ch1fp Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Data Sheet For Ch1fp
 - Highlighting and Note-Taking Chemistry Data Sheet For Ch1fp
 - Interactive Elements Chemistry Data Sheet For Ch1fp
- 8. Staying Engaged with Chemistry Data Sheet For Ch1fp
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Data Sheet For Ch1fp
- 9. Balancing eBooks and Physical Books Chemistry Data Sheet For Ch1fp
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Data Sheet For Ch1fp
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Data Sheet For Ch1fp
 - Setting Reading Goals Chemistry Data Sheet For Ch1fp
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Data Sheet For Ch1fp
 - Fact-Checking eBook Content of Chemistry Data Sheet For Ch1fp
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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