

NAME: \_\_\_\_\_ Period: \_\_\_\_\_

Ch 10 Chemical Quantities

Study Guide

/75

Matching

Match each item with the correct statement below.

- a. molar volume
- b. molar mass
- c. atomic mass

- \_\_\_\_\_ 1. the number of grams of an element that is numerically equal to the atomic mass of the element in amu
- \_\_\_\_\_ 2. the mass of a mole of any element or compound
- \_\_\_\_\_ 3. the volume occupied by a mole of any gas at STP

Match each item with the correct statement below.

- a. representative particle
- b. mole
- c. Avogadro's number
- d. percent composition
- e. standard temperature and pressure
- f. empirical formula

- \_\_\_\_\_ 4. the number of representative particles of a substance present in 1 mole of that substance
- \_\_\_\_\_ 5. an atom, an ion, or a molecule, depending upon the way a substance commonly exists
- \_\_\_\_\_ 6. the SI unit used to measure amount of substance
- \_\_\_\_\_ 7. 0°C and 1 atm
- \_\_\_\_\_ 8. the percent by mass of each element in a compound
- \_\_\_\_\_ 9. the smallest whole number ratio of the atoms in a compound

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 1. What SI unit is used to measure the number of representative particles in a substance?  
a. kilogram      b. ampere      c. kelvin      d. mole
- \_\_\_\_\_ 2. How many hydrogen atoms are in 5 molecules of isopropyl alcohol,  $C_3H_7O$ ?  
a.  $5 \times (6.02 \times 10^{23})$       c. 35  
b. 5      d.  $35 \times (6.02 \times 10^{23})$
- \_\_\_\_\_ 3. If a dozen Clementine oranges have a mass of 744 g, what will be the mass of 15 Clementines?  
a. 11.16 kg      b. 595 g      c. 896 g      d. 930 g
- \_\_\_\_\_ 4. Which of the following elements exists as a diatomic molecule?  
a. neon      b. lithium      c. nitrogen      d. sulfur
- \_\_\_\_\_ 5. Avogadro's number of representative particles is equal to one \_\_\_\_\_.  
a. kilogram      b. gram      c. kelvin      d. mole
- \_\_\_\_\_ 6. Which of the following is equal to Avogadro's number?  
a. the number of atoms of bromine in 1 mol  $Br_2$   
b. the number of molecules of chlorine in 0.50 mol  $Cl_2$   
c. the number of molecules of nitrogen in 1 mol  $N_2$   
d. the number of atoms in 1 mol CO

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**Siddappa N.Byrareddy**



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## **Table of Contents Chemistry Chemical Quantities Study Guide Answers**

1. Understanding the eBook Chemistry Chemical Quantities Study Guide Answers
  - The Rise of Digital Reading Chemistry Chemical Quantities Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chemical Quantities Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chemical Quantities Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chemical Quantities Study Guide Answers

- Personalized Recommendations
  - Chemistry Chemical Quantities Study Guide Answers User Reviews and Ratings
  - Chemistry Chemical Quantities Study Guide Answers and Bestseller Lists
5. Accessing Chemistry Chemical Quantities Study Guide Answers Free and Paid eBooks
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    - Chemistry Chemical Quantities Study Guide Answers eBook Subscription Services
    - Chemistry Chemical Quantities Study Guide Answers Budget-Friendly Options
  6. Navigating Chemistry Chemical Quantities Study Guide Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Chemistry Chemical Quantities Study Guide Answers Compatibility with Devices
    - Chemistry Chemical Quantities Study Guide Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemistry Chemical Quantities Study Guide Answers
    - Highlighting and Note-Taking Chemistry Chemical Quantities Study Guide Answers
    - Interactive Elements Chemistry Chemical Quantities Study Guide Answers
  8. Staying Engaged with Chemistry Chemical Quantities Study Guide Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Chemical Quantities Study Guide Answers
  9. Balancing eBooks and Physical Books Chemistry Chemical Quantities Study Guide Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Chemical Quantities Study Guide Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Chemical Quantities Study Guide Answers
    - Setting Reading Goals Chemistry Chemical Quantities Study Guide Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Chemical Quantities Study Guide Answers

- Fact-Checking eBook Content of Chemistry Chemical Quantities Study Guide Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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