

General Chemistry II Chapter 15 Study Guide

- $S_{\text{gas}} > S_{\text{liquid}} > S_{\text{solid}}$
- For similar molecules, S° tends to increase with:
 - Increasing molar mass
 - Increased number of atoms in the formula

Standard entropy of a reaction ($\Delta S_{\text{rxn}}^\circ$):

- $\Delta S_{\text{rxn}}^\circ$: is the entropy change that occurs when all reactants and products are in their standard states:

$$\Delta S_{\text{rxn}}^\circ = \sum m S_{\text{products}}^\circ - \sum n S_{\text{reactants}}^\circ$$

where m and n are the mols of products and reactants, given by the coefficients in the balanced equation.

Suppose a reaction: $a \text{ A} + b \text{ B} \rightarrow c \text{ C} + d \text{ D}$

$$\sum m S_{\text{products}}^\circ = c \Delta S^\circ (\text{C}) + d \Delta S^\circ (\text{D})$$

$$\sum n S_{\text{reactants}}^\circ = a \Delta S^\circ (\text{A}) + b \Delta S^\circ (\text{B})$$

Remember to multiply ΔH° of each substance by its coefficient.

Predicting the sign of $\Delta S_{\text{rxn}}^\circ$:

- We can often predict the sign of $\Delta S_{\text{rxn}}^\circ$ for processes that involve a change in the number of moles of gas
 - $\Delta S_{\text{rxn}}^\circ$ is **positive** if the amount of gas increases
 - $\Delta S_{\text{rxn}}^\circ$ is **negative** if the amount of gas decreases
 - $\Delta S_{\text{rxn}}^\circ$ is **likely to be positive** if a new structure forms that has more freedom of motion

Gibbs free energy:

- Gibbs free energy (G): combines the enthalpy and entropy of a system:
 - $G = H - TS$
- Gibbs free energy change (ΔG): is a measure of the spontaneity of a process and of the useful energy available from it:
 - $\Delta G_{\text{rxn}} = \Delta H_{\text{rxn}} - T\Delta S_{\text{rxn}}$
- $\Delta G < 0$ for a spontaneous reaction ($-\Delta G$)
- $\Delta G > 0$ for a nonspontaneous reaction ($+\Delta G$)
- $\Delta G = 0$ for a process at equilibrium ($\Delta G = 0$)

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