

10

CHEMICAL QUANTITIES

Practice Problems

On a piece of paper, solve the following problems.

YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER

1. What is the molar mass of sucrose ($C_{12}H_{22}O_{11}$)?
2. What is the molar mass of each of the following compounds?
 - a. phosphorus pentachloride (PCl_5)
 - b. uranium hexafluoride (UF_6)
3. Calculate the molar mass of each of the following ionic compounds:
 - a. $KMnO_4$
 - b. $Ca_3(PO_4)_2$
4. How many moles is 3.52×10^{24} molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen (O_2)?

SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS

1. What is the molar mass of each of the following compounds?
 - a. $C_6H_{12}O_6$
 - b. $NaHCO_3$
 - c. C_2H_{12}
 - d. KNH_4SO_4
2. Calculate the mass in grams of each of the following:
 - a. 8.0 mol lead oxide (PbO)
 - b. 0.75 mol hydrogen sulfide (H_2S)
 - c. 0.00100 mol silicon tetrahydride (SiH_4)
 - d. 1.50×10^{-2} mol molecular oxygen (O_2)
 - e. 2.30 mol ethylene glycol ($C_2H_6O_2$)
3. How many grams are in 1.73 mol of dinitrogen pentoxide (N_2O_5)?
4. How many grams are in 0.658 mol of calcium phosphate [$Ca_3(PO_4)_2$]?
5. Calculate the number of moles in each of the following:
 - a. 0.50 g sodium bromide ($NaBr$)
 - b. 13.5 g magnesium nitrate [$Mg(NO_3)_2$]
 - c. 1.02 g magnesium chloride ($MgCl_2$)
 - d. 0.00100 g monochloromethane (CH_3Cl)
 - e. 1.50×10^{-3} g propylene glycol [$C_3H_8(OH)_2$]
6. A chemist plans to use 435.0 grams of ammonium nitrate (NH_4NO_3) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ($C_{20}H_{24}N_2O_2$). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane (CH_4) gas?
9. How many moles is 135 L of ammonia (NH_3) gas at STP?

Chemistry Chapter 10 Study Guide

Mark M. Green, Harold A. Wittcoff



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principles that are often taught to students without reference to the commercial and practical importance of the subject. The marvelous theoretical principles on which organic chemistry is based are therefore often not fully appreciated by students of this subject. Organic chemistry can appear dry, meaningless, and seemingly irrelevant. In this textbook, designed to be used in conjunction with classic texts of organic chemistry at the undergraduate level or standing alone for more advanced students, two experts, M. M. Green and H. A. Wittcoff, bring the principles and the practice together. Written for students and also giving much information that could be used to enhance teaching of the subject, the book presented in ten concise chapters combines important industrial processes with the principles of organic chemistry. The result is a source of otherwise barely accessible information. In addition, personal anecdotes from the authors' vast experience make this a fascinating and indispensable textbook for everyone wishing to enhance the appreciation of this subject. I have never come across such an enticing mix of stories of discovery with basic chemistry. Roald Hoffmann, Cornell University. Simply put, this book is a gem. The chemistry described is rigorous, but the warm, humorous, and conversational writing style makes the book a joy to read. Dasan M. Thamattoor, Colby College. This is a unique, fascinating book that bridges organic chemistry principles with chemical industrial applications. The story-telling style makes the reading/learning experience extremely enjoyable. Qiao Sheng Hu, College of Staten Island, City University of New York.

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