

8. Complete the following chart.

	Pure Substance or Mixture?	Element, Compound, Homogeneous, or Heterogeneous?
Air	Mixture	Homogeneous
Carbon (C)	Pure Substance	Element
Water (H ₂ O)	Pure Substance	Compound
Chocolate-Chip Cookies	Mixture	Heterogeneous

Unit 2: The Atom

Important Vocabulary & their Definitions:

- Proton: positive atomic particles
- Neutron: neutral atomic particles
- Electron: negative atomic particles
- Isotopes: same element with different numbers of neutrons

1. Use the element tile below to determine the number of protons, neutrons, and electrons in an atom of this element.

Protons = 13
Neutrons = 14
Electrons = 13



2. Calculate the average atomic mass given the chart below.

Isotope	²⁷ Al	²⁸ Al	²⁹ Al
Atomic Mass	26.9815 amu	27.9819 amu	28.9819 amu
% Abundance	90.62%	7.83%	1.55%

$$15.98 \times 0.9062 + 27.98 \times 0.0783 + 28.98 \times 0.0155 = 26.98$$

Unit 3: The Periodic Table

Important Vocabulary & their Definitions:

- Valence Electrons: electrons in outer most shell
- Octet Rule: all atoms want a full valence shell
- Atomic Radius: distance from nucleus to outermost shell
- Electronegativity: ability to pull electrons
- Ionization Energy: energy required to remove an electron

Important Scientists & why they are important:

- Mendeleev: put elements in groups based on properties, left blanks for undiscovered elements

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Michael Brown



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Hazmat Chemistry Study Guide (Second Edition) Jill Meryl Levy, 2011

Chemical Principles Study Guide/Solutions

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