

CHEM 110
Extra In-Class Practice Problems
Lecture Set 4

1. Molarity

- A. 7.93 g sugar ($C_{12}H_{22}O_{11}$) is dissolved in water such that the final volume of the solution is 750. mL. Calculate the concentration (M).

$$M = \frac{\text{mol}}{L} = \frac{0.02316 \text{ mol } C_{12}H_{22}O_{11}}{0.750 L} = \boxed{0.0309 M} \quad 3SF$$

$$\text{moles } C_{12}H_{22}O_{11}: \frac{7.93 g}{342.34 g} = 0.02316 \text{ mol } C_{12}H_{22}O_{11} \quad 3SF$$

- B. What is the molarity of a solution made by dissolving 0.075 mol of NaCl in 25.0 mL of solution?

$$M = \frac{\text{mol}}{L} = \frac{0.075 \text{ mol NaCl}}{0.0250 L} = \boxed{3.0 M} \quad 2SF$$

- C. A solution is made by dissolving 18.9 g KCl in water and diluting to 500. mL. What is the molarity of this solution?

$$M = \frac{\text{mol}}{L} = \frac{0.2535 \text{ mol KCl}}{0.500 L} = \boxed{0.507 M} \quad 3SF$$

$$\text{moles KCl}: \frac{18.9 g}{74.55 g} = 0.2535 \text{ mol KCl}$$

- D. How many moles of HCl are contained in 75.0 mL of 1.85 M HCl?

$$\frac{0.0750 L}{1 L} \times \frac{1.85 \text{ mol HCl}}{1 L} = \boxed{0.139 \text{ mol HCl}} \quad 3SF$$

\Downarrow
 $\frac{1.85 \text{ mol HCl}}{1 L} \}$ conversion

Chemistry 110 Practice Problems

Yicheng Fang



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Chemistry 110 Practice Problems Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has been much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Chemistry 110 Practice Problems**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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