

Multiple Choice: *Identify the letter of the choice that best completes the statement or answers the question.*

- ____ 1. What scientific (natural) law serves as the basis for balancing chemical equations by requiring that there be no observable change in the quantity of matter during a chemical reaction?
- Law of Conservation of Energy
 - Law of Multiple Proportions
 - Law of Conservation of Matter
 - Law of Definite Proportions
 - Law of Constant Composition
- ____ 2. Balance the following equation with the **smallest whole number coefficients**. What is the coefficient for O_2 in the balanced equation?



- 9
 - 5
 - 15
 - 6
 - 13
- ____ 3. What is the coefficient for HBr when the following equation is balanced with the **smallest whole number coefficients**?
- $$Br_2 + H_2O \rightarrow HBr + HBrO_3$$
- 5
 - 7
 - 8
 - 3
 - 6
- ____ 4. Balance the following equation with the **smallest whole number coefficients**. What is the coefficient for NH_3 in the balanced equation?
- $$Fe(NO_3)_3 + NH_3 + H_2O \rightarrow Fe(OH)_3 + NH_4NO_3$$
- 1
 - 3
 - 2
 - 6
 - 4
- ____ 5. Elemental phosphorus is produced from calcium phosphate in the following reaction. What is the coefficient for C when this equation is balanced with the **smallest whole number coefficients**?
- $$Ca_3(PO_4)_2 + SiO_2 + C \rightarrow P_4 + CO + CaSiO_3$$

- 10
- 3
- 1
- 6
- 4

Chemistry Semester 2 Final Form A Answers

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