

**THE FOLLOWING QUESTIONS ARE PRACTICE QUESTIONS FOR THE CSN CHEMISTRY PLACEMENT EXAM**

The chemistry placement test is used to assess your preparation for General Chemistry I, CHEM 121 at CSN. The test is a standardized test that will examine your knowledge of chemistry and basic mathematical skills. You will be provided scratch paper, a Periodic Table of the Elements and a nonprogrammable scientific calculator. The following sample questions are intended only as a guide to show you the relative level of difficulty and example topics covered in the Chemistry Placement Exam. The actual exam will have a time limit of 45 minutes to answer 44 multiple choice questions. Click the following link to view a [Periodic Table](#).

**MULTIPLE CHOICE.** Choose the single choice that best answers the question.

- 1) How many total atoms are in the formula



A) 12                      B) 14                      C) 9                      D) 8

- 2) How many electrons are in  $\text{Br}^-$ ?

A) 4                      B) 7                      C) 36                      D) 34

- 3) What are the coefficients for the following reaction when it is properly balanced?



A) 2, 3, 6, 1                      B) 6, 1, 3, 2  
C) 2, 3, 1, 6                      D) 2, 1, 1, 3

- 4) Given that  $4 \text{ NH}_3 + 5 \text{ O}_2 \rightarrow 4 \text{ NO} + 6 \text{ H}_2\text{O}$ , when 4.50 mol of  $\text{H}_2\text{O}$  are formed, the amount of NO formed is

A) 3.00 mol                      B) 6.75 mol  
C) 1.50 mol                      D) 4.50 mol

- 5) What type of a reaction occurs when a sodium hydroxide solution is mixed with an acetic acid solution?

A) gas evolution  
B) oxidation-reduction  
C) acid-base neutralization  
D) precipitation

- 6) Which of the following atoms has the smallest radius?

A) Ne                      B) O  
C) Li                      D) B

- 7) What is the molecular geometry if you have a double bond, a single bond and 1 lone pair around the central atom?

A) trigonal pyramidal  
B) linear  
C) tetrahedral  
D) bent

- 8) What volume of 12.0 M HCl is required to make 75.0 mL of 3.50 M HCl?

A) 0.560 mL                      B) 257 mL  
C) 560. mL                      D) 21.9 mL

$$9) \frac{(2.80 \times 10^{-3})(1.28 \times 10^1)}{(3.20 \times 10^{-4})} = \underline{\hspace{2cm}}$$

A)  $1.27 \times 10^1$                       B)  $2.0 \times 10^2$   
C)  $1.2 \times 10^6$                       D)  $1.12 \times 10^2$

- 10) Of the molecules below, only \_\_\_\_\_ is polar.

A)  $\text{CH}_4$                       B)  $\text{SbF}_5$   
C)  $\text{CO}_2$                       D)  $\text{AsH}_3$

- 11) Which one of the following is a metalloid?

A) S                      B) Ge                      C) C                      D) Pb

- 12) A certain atom has a charge of -2 and has 16 protons and 16 neutrons in the nucleus. How many electrons does this atom possess?

A) 34                      B) 14                      C) 32                      D) 18

- 13) How many moles of pyridine ( $\text{C}_5\text{H}_5\text{N}$ ) are contained in 3.13 g of pyridine?

A) 0.00404                      B) 25.3  
C) 0.0396                      D) 0.319

- 14) Solve the following equation for x:

$$4x - b = 2x + c$$

A)  $c - b - 6$                       B)  $\frac{c + b}{2}$   
C)  $\frac{c + b}{6}$                       D)  $\frac{c - b}{2}$

# Chemistry Readiness Test Math Sample

**Clemens Wendtner**



## **Chemistry Readiness Test Math Sample:**

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**Material Tested** The following topics are tested on the SAT II Math Level IC

**Artificial Intelligence in the Clinical Laboratory: Current Practice and Emerging Opportunities, An Issue of the Clinics in Laboratory Medicine, E-Book** Jason Baron, 2023-02-10 In this issue guest editors bring their considerable expertise to this important topic Provides in depth reviews on the latest updates in the field providing actionable insights for clinical practice Presents the latest information on this timely focused topic under the leadership of experienced editors in the field Authors synthesize

**Educational Assessment** Karlheinz Ingenkamp, 1977

**Resources in Education**, 1999

**Competition Science Vision**, 2000-07

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