



# CHEMISTRY

## CHEAT SHEET



# INTERMOLECULAR FORCES

## INTERMOLECULAR FORCES CHEAT SHEET

Intermolecular forces are attractive forces between molecules. They are responsible for many physical properties and chemical behavior of substances.

The strength of intermolecular forces depends on the nature of the molecules. The order of intermolecular forces is as follows: Hydrogen bonding > Dipole-dipole interaction > London dispersion forces.

Intermolecular Force	Description	Types of Molecules
Hydrogen Bonding	Occurs when a hydrogen atom is bonded to a highly electronegative atom (F, O, N). The electronegative atom attracts the hydrogen atom, creating a partial positive charge on the hydrogen atom. This partial positive charge is attracted to the partial negative charge on the electronegative atom of another molecule.	Small molecules containing F, O, or N
Dipole-Dipole Interaction	Occurs between polar molecules. Each molecule has a partial positive end and a partial negative end. The partial positive end of one molecule is attracted to the partial negative end of another molecule.	Polar molecules
London Dispersion Forces	Occurs between all molecules, regardless of whether they are polar or non-polar. These forces are caused by the temporary dipoles that form in all molecules.	All molecules

Intermolecular forces are responsible for many physical properties of substances, such as boiling point, melting point, and viscosity. The strength of intermolecular forces increases with the size and complexity of the molecules.

**Hydrogen Bonding**

Hydrogen bonding is a strong intermolecular force that occurs between molecules containing hydrogen atoms bonded to highly electronegative atoms (F, O, N).

**Dipole-Dipole Interaction**

Dipole-dipole interaction is a moderate intermolecular force that occurs between polar molecules.

**London Dispersion Forces**

London dispersion forces are weak intermolecular forces that occur between all molecules, regardless of whether they are polar or non-polar.

The strength of intermolecular forces increases with the size and complexity of the molecules. The following table shows the boiling and melting points of several substances, which are indicators of the strength of their intermolecular forces.

Substance	Formula	Boiling Point (°C)	Melting Point (°C)
Water	$H_2O$	100	0
Alcohol	$C_2H_5OH$	78	-114
Gasoline	$C_8H_{18}$	125	-95

As you can see, the boiling and melting points of water are much higher than those of alcohol and gasoline, indicating that water has the strongest intermolecular forces.

# Chemistry Physics Forces Study Guide

**K Morrison**



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