

Ionic Compounds and Metals

Section 7.1 Ion Formation

pages 206–209

Section 7.1 Assessment

page 209

1. **Compare** the stability of a lithium atom with that of its ion, Li^+ .

The Li^+ ion is more stable because it has a complete octet.

2. **Describe** two different causes of the force of attraction in a chemical bond.

The two causes are the attraction between the positive nucleus of one atom and the negative electrons of another atom, and the attraction between positive ions and negative ions.

3. **Apply** Why are all of the elements from group 18 relatively unreactive, whereas those in group 17 are very reactive?

The group 18 elements, known as noble gases, have complete outer energy levels and do not easily form ions. The group 17 elements are highly reactive because they need only to gain a single electron to form an octet.

4. **Summarize** ionic bond formation by correctly pairing these terms: *cation*, *anion*, *electron gain*, *electron loss*.

(anion, electron gain), (cation, electron loss)

5. **Apply** Write out the electron configuration for each atom. Then predict the change that must occur in each to achieve a noble gas configuration.

- a. nitrogen

$[\text{He}]2s^22p^3$; gain 3 electrons (3^- ion) or lose 5 electrons (5^+ ion)

- b. sulfur

$[\text{Ne}]3s^23p^4$; gain 2 electrons (2^- ion)

- c. barium

$[\text{Xe}]6s^2$; lose 2 electrons (2^+ ion)

- d. lithium

$[\text{He}]2s^1$; lose 1 electron (1^+ ion)

6. **Model** Draw models to represent the formation of the positive calcium ion and the negative bromide ion.

Models should show that the calcium atom loses two electrons, forming Ca^{2+} , and that bromine gains one electron, forming Br^- . The models should also show the addition of energy to form the calcium ion, and the release of energy to form the bromide ion.

Section 7.2 Ionic Bonds and Ionic Compounds

pages 210–217

Practice Problems

page 212

Explain how an ionic compound forms from these elements.

7. sodium and nitrogen

Three Na atoms each lose $1 e^-$, forming 1^+ ions. One N atom gains $3 e^-$, forming a 3^- ion. The ions attract, forming Na_3N .

$$3 \text{ Na ions } \left(\frac{1+}{\text{Na ion}} \right) + 1 \text{ N ion } \left(\frac{3-}{\text{N ion}} \right) \\ = 3(1+) + 1(3-) = 0$$

The overall charge on one formula unit of Na_3N is zero.

8. lithium and oxygen

Two Li atoms each lose $1 e^-$, forming 1^+ ions. One O atom gains $2 e^-$, forming a 2^- ion. The ions attract, forming Li_2O .

$$2 \text{ Li ions } \left(\frac{1+}{\text{Li ion}} \right) + 1 \text{ O ion } \left(\frac{2-}{\text{O ion}} \right) \\ = 2(1+) + 1(2-) = 0$$

The overall charge on one formula unit of Li_2O is zero.

Chemistry Matter And Change Chapter 7 Study Guide Answers

James E. Brady, Fred Senese



Chemistry Matter And Change Chapter 7 Study Guide Answers:

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Basic Concepts of Chemistry, Study Guide Leo J. Malone, 2000 The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems

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numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students Study Guide with Answers to Selected Problems H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises *Study Guide and Solutions Manual* Neil E. Schore,2002-08-02 **Ebook:** **Chemistry: The Molecular Nature of Matter and Change** Silberberg,2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change **Know Your 'O' Level Chemistry - A Study Guide** , *Study Guide and Full Solutions Manual* Susan McMurry,John McMurry,2002-08 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all end of chapter problems in the textbook **Books and Pamphlets, Including Serials and Contributions to Periodicals** Library of Congress. Copyright Office,1968 **Life: The Science of Biology Study Guide** William K. Purves,Edward Dzialowski,Lindsay Goodloe,Betty McGuire,Nancy Guild,Paula Mabee,2003-12-26 New edition of a text presenting underlying concepts and showing their relevance to medical agricultural and environmental issues Seven chapters discuss the cell information and heredity evolutionary process the evolution of diversity the biology of flowering plants and of animals and ecology and biogeography Topics are linked by themes such as evolution the experimental foundations of knowledge the flow of energy in the living world the application and influence of molecular techniques and human health considerations Includes a CD ROM which covers some of the subject matter and introduces and illustrates 1 700 plus key terms and concepts Annotation copyrighted by Book News Inc Portland OR Resources for Teaching Middle School Science Smithsonian Institution,National Academy of Engineering,National Science Resources Center of the National Academy of Sciences,Institute of Medicine,1998-03-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them Resources for Teaching Middle School Science developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the National Science Education Standards This completely new guide follows on the success of Resources for Teaching Elementary School Science the first in the NSRC series of annotated guides to hands on inquiry centered curriculum materials and other resources for science teachers The curriculum materials in the new guide are grouped in five chapters by

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