

## CHEMISTRY FINAL EXAM REVIEW GUIDE 2012 Scott

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### Multiple Choice

*Identify the choice that best completes the statement or answers the question.*

- \_\_\_\_\_ 1) Which of the following observations is quantitative?  
a. The liquid turns blue litmus paper red.  
b. The liquid boils at 100°C.  
c. The liquid tastes bitter.  
d. The liquid is cloudy.
- \_\_\_\_\_ 2) The symbols for units of length in order from largest to smallest are  
a. m, cm, mm, km. c. km, mm, cm, m.  
b. mm, m, cm, km. d. km, m, cm, mm.
- \_\_\_\_\_ 3) A volume of 1 milliliter is equivalent to  
a. 1 cubic centimeter. c. 1 liter.  
b. 1 gram. d.  $10^{-3}$  cubic decimeters.
- \_\_\_\_\_ 4) Which pair of quantities determines the density of a material?  
a. mass and weight c. volume and concentration  
b. volume and weight d. volume and mass
- \_\_\_\_\_ 5) A sample of gold has a mass of 96.5 g and a volume of 5.00 cm<sup>3</sup>. The density of gold is  
a. 0.0518 g/cm<sup>3</sup>. c. 101.5 g/cm<sup>3</sup>.  
b. 19.3 g/cm<sup>3</sup>. d. 483 g/cm<sup>3</sup>.
- \_\_\_\_\_ 6) The number of grams equal to 0.5 kg is  
a. 0.0005. c. 500.  
b. 0.005. d. 5000.
- \_\_\_\_\_ 7) If 1 inch equals 2.54 cm, how many centimeters equal 1 yard?  
a. 0.0706 cm c. 30.5 cm  
b. 14.2 cm d. 91.4 cm
- \_\_\_\_\_ 8) A chemist obtained the values 5.2246 g, 5.2353 g, and 5.2501 g for the mass of a sample. Without knowing the true mass of the sample, it can be said that these values have  
a. good precision.  
b. good accuracy.  
c. poor precision.  
d. poor accuracy.
- \_\_\_\_\_ 9) The measurement 0.035550 g rounded off to two significant figures would be  
a. 0.03 g. c. 0.036 g.  
b. 0.35 g. d.  $3.5 \times 10^2$  g.
- \_\_\_\_\_ 10) When 64.4 is divided by 2.00, the correct number of significant figures in the result is

# Chemistry Final Exam Study Guide Problems

**SAT II Exambusters**



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*Hazmat Chemistry Study Guide (Second Edition)* Jill Meryl Levy, 2011

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**CliffsNotes AP Chemistry** Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book

[Study Guide for Chemistry by Steven S. Zumdahl](#) Martha B. Barrett, 1986

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**Making it relevant** Peter Nentwig,David Waddington,2006 Teaching in context has become an accepted and often welcomed way of teaching science in both primary and secondary schools The conference organised by IPN and the University of York Science Education Group Context based science curricula drew on the experience of over 40 science educators and 10 projects The book is arranged in four parts Part A consists of two papers one on situated learning and the other on implementation of new curricula Part B contains descriptions of five major curricula in different countries why they were introduced how they were developed and implemented and evaluation results Part C gives descriptions of three projects that are of smaller scale and their materials are used as interventions in other more conventional curricula There is also a contribution on some fundamental research where modules of work are written to examine how best to design context based curricula Finally Part D consist of two chapters one summarising some of the findings that came out of the chapters in the three earlier parts and the second looks at the future

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*The Hidden Curriculum—Faculty-Made Tests in Science* Sheila Tobias,Jacqueline Raphael,2013-06-29 This resource manual for college level science instructors reevaluates the role of testing in their curricula and describes innovative techniques pioneered by

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web we utilize the equation that relate amount remaining initial mass and number of half lives  $n = \frac{t}{t_{1/2}} \times \ln 2$   
 125 g 5 what is the half life of an isotope that is 80 remained after 16 days solution remaining 80 100 therefore  $n = \frac{t}{t_{1/2}} \times \ln 2$   
 now using the half life equation in example 1 we have

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solution 20.0  $\div$  3.6 = 5.56 half lives

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