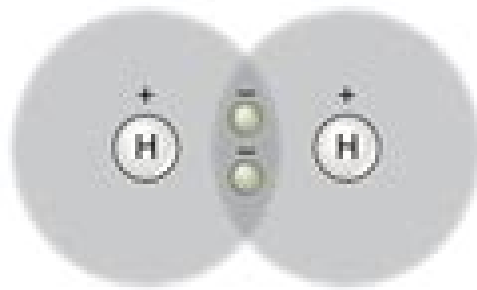
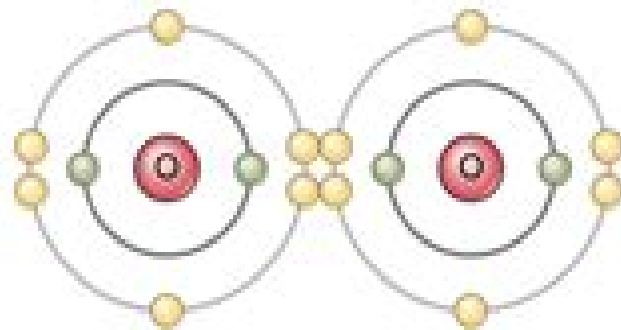


(a) A single covalent bond: hydrogen gas ($\text{H}-\text{H}$). Two atoms of hydrogen each share their solitary electron in a single covalent bond.

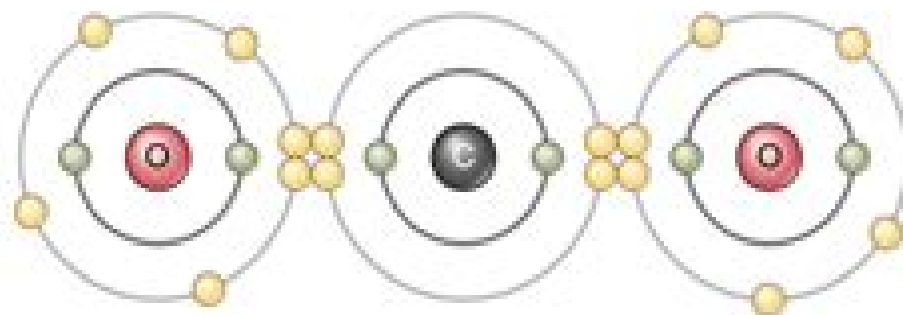


(b) A double covalent bond: oxygen gas ($\text{O}=\text{O}$). An atom of oxygen has six electrons in its valence shell; thus, two more would make it stable. Two atoms of oxygen achieve stability by sharing two pairs of electrons in a double covalent bond.



Molecule of oxygen gas (O_2)

(c) Two double covalent bonds: carbon dioxide ($\text{O}=\text{C}=\text{O}$). An atom of carbon has four electrons in its valence shell; thus, four more would make it stable. An atom of carbon and two atoms of oxygen achieve stability by sharing two electron pairs each, in two double covalent bonds.



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