



# CHEMISTRY

## CHEAT SHEET

# EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol  $\rightleftharpoons$ .

<p><b>What is Equilibrium?</b></p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.</p>	<p><b>Equilibrium Constants</b></p> <p><math>K_c</math> (Concentration): <math>K_c = \frac{[Products]}{[Reactants]}</math></p> <p><math>K_p</math> (Pressure): <math>K_p = \frac{P_{Products}}{P_{Reactants}}</math></p> <p><math>K_a</math> (Acid Dissociation): <math>K_a = \frac{[H^+][A^-]}{[HA]}</math></p> <p><math>K_b</math> (Base Dissociation): <math>K_b = \frac{[OH^-][BH^+]}{[B]}</math></p>											
<p><b>Equilibrium Constant, <math>K_c</math></b></p> <p>For a general reaction: <math>aA + bB \rightleftharpoons cC + dD</math></p> <p>The equilibrium constant <math>K_c</math> is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ <p>Where <math>[A]</math>, <math>[B]</math>, <math>[C]</math>, and <math>[D]</math> are the concentrations of the reactants and products at equilibrium.</p>	<table><tr><th>Equilibrium Constant, <math>K_p</math></th><th>Meaning</th></tr><tr><td><math>K_p &gt; 1</math></td><td>Products are favored at equilibrium.</td></tr><tr><td><math>K_p &lt; 1</math></td><td>Reactants are favored at equilibrium.</td></tr><tr><td><math>K_p = 1</math></td><td>Reactants and products are at equal concentrations.</td></tr></table>	Equilibrium Constant, $K_p$	Meaning	$K_p > 1$	Products are favored at equilibrium.	$K_p < 1$	Reactants are favored at equilibrium.	$K_p = 1$	Reactants and products are at equal concentrations.			
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<p><b>Equilibrium Constant, <math>K_a</math></b></p> <p>For a weak acid <math>HA</math> dissociating in water:</p> $HA \rightleftharpoons H^+ + A^-$ <p>The equilibrium constant <math>K_a</math> is given by:</p> $K_a = \frac{[H^+][A^-]}{[HA]}$	<table><tr><th>Equilibrium Constant, <math>K_b</math></th><th>Meaning</th></tr><tr><td><math>K_b &gt; 1</math></td><td>Products are favored at equilibrium.</td></tr><tr><td><math>K_b &lt; 1</math></td><td>Reactants are favored at equilibrium.</td></tr><tr><td><math>K_b = 1</math></td><td>Reactants and products are at equal concentrations.</td></tr></table>	Equilibrium Constant, $K_b$	Meaning	$K_b > 1$	Products are favored at equilibrium.	$K_b < 1$	Reactants are favored at equilibrium.	$K_b = 1$	Reactants and products are at equal concentrations.			
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<p><b>Le Chatelier's Principle: The effect of changing conditions on equilibrium</b></p>												
<table><tr><th>Condition</th><th>Effect on Equilibrium</th><th>Example</th></tr><tr><td>Concentration</td><td><ul style="list-style-type: none"><li>Increasing the concentration of a reactant shifts the equilibrium to the right (towards the products).</li><li>Increasing the concentration of a product shifts the equilibrium to the left (towards the reactants).</li></ul></td><td><math display="block">N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)</math><p>Increasing the concentration of <math>N_2</math> or <math>H_2</math> will shift the equilibrium to the right, producing more <math>NH_3</math>.</p></td></tr><tr><td>Pressure</td><td><ul style="list-style-type: none"><li>Increasing the pressure (by decreasing the volume) shifts the equilibrium towards the side with fewer moles of gas.</li><li>Decreasing the pressure (by increasing the volume) shifts the equilibrium towards the side with more moles of gas.</li></ul></td><td><math display="block">N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)</math><p>Increasing the pressure will shift the equilibrium to the right, as there are fewer moles of gas on the product side (2 moles) compared to the reactant side (4 moles).</p></td></tr><tr><td>Temperature</td><td><ul style="list-style-type: none"><li>For an exothermic reaction, increasing the temperature shifts the equilibrium to the left (towards the reactants).</li><li>For an endothermic reaction, increasing the temperature shifts the equilibrium to the right (towards the products).</li></ul></td><td><math display="block">N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) \quad \Delta H = -92 \text{ kJ/mol}</math><p>Since the reaction is exothermic, increasing the temperature will shift the equilibrium to the left, producing less <math>NH_3</math>.</p></td></tr></table>	Condition	Effect on Equilibrium	Example	Concentration	<ul style="list-style-type: none"><li>Increasing the concentration of a reactant shifts the equilibrium to the right (towards the products).</li><li>Increasing the concentration of a product shifts the equilibrium to the left (towards the reactants).</li></ul>	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ <p>Increasing the concentration of <math>N_2</math> or <math>H_2</math> will shift the equilibrium to the right, producing more <math>NH_3</math>.</p>	Pressure	<ul style="list-style-type: none"><li>Increasing the pressure (by decreasing the volume) shifts the equilibrium towards the side with fewer moles of gas.</li><li>Decreasing the pressure (by increasing the volume) shifts the equilibrium towards the side with more moles of gas.</li></ul>	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ <p>Increasing the pressure will shift the equilibrium to the right, as there are fewer moles of gas on the product side (2 moles) compared to the reactant side (4 moles).</p>	Temperature	<ul style="list-style-type: none"><li>For an exothermic reaction, increasing the temperature shifts the equilibrium to the left (towards the reactants).</li><li>For an endothermic reaction, increasing the temperature shifts the equilibrium to the right (towards the products).</li></ul>	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) \quad \Delta H = -92 \text{ kJ/mol}$ <p>Since the reaction is exothermic, increasing the temperature will shift the equilibrium to the left, producing less <math>NH_3</math>.</p>
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# Chemistry Chemical Equilibrium Answers Study Guide

**Minjie Lin**



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**Chemistry, Study Guide** Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460

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