

**Chapter 12 Stoichiometry Review**

**Directions: Show all work!!! No Work Shown, no credit given! Include all units and chemical formulas.**

- Balance the following equation and work the problems below:  $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$ 
  - If 6.52 L of  $\text{H}_2$  react with excess nitrogen what volume of  $\text{NH}_3$  is produced?
  - In order to produce 6.52 L of  $\text{NH}_3$  how many liters of nitrogen are needed?
  - If  $2.35 \times 10^{24}$  molecules of  $\text{NH}_3$  is formed how many grams of hydrogen was used?
- Use the following equation answer these questions:  $\text{Mg} + 2 \text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + \text{H}_2$ 
  - How many grams of magnesium is need to react with 6.28 grams  $\text{HNO}_3$ ?
  - If I start this reaction with 40.5 grams of magnesium and an excess of nitric acid,
    - how many grams of hydrogen gas will I produce?
    - If 1.7 grams of hydrogen is actually produced, what was my percent yield of hydrogen?
- Use the following equation answer these questions:  $\text{NaHCO}_3 \rightarrow \text{NaOH} + \text{CO}_2$ 
  - Find the theoretical yield in grams of  $\text{NaOH}$  if 14.50 grams  $\text{NaHCO}_3$  is decomposed.
  - If 25 liters of carbon dioxide gas is produced in this reaction,
    - how many grams of sodium hydroxide should be produced?
    - If 50 grams of sodium hydroxide are actually produced, what was my percent yield?
- Use the following equation answer these questions:  $2 \text{FePO}_4 + 3 \text{Na}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 2 \text{Na}_3\text{PO}_4$ 
  - How many grams of  $\text{FePO}_4$  is needed to produce  $2.67 \times 10^{25}$   $\text{Na}_3\text{PO}_4$  formula units
  - If I start with 25 g of  $\text{FePO}_4$ 
    - how many grams of  $\text{Fe}_2(\text{SO}_4)_3$  can I make?
    - If 18.5 grams of  $\text{Fe}_2(\text{SO}_4)_3$  are actually made what is the percent yield?
- Use the following equation answer these questions:  $\text{C}_{12}\text{H}_{22}\text{O}_{11} + 12\text{O}_2 \rightarrow 12\text{CO}_2 + 11\text{H}_2\text{O}$ 
  - If 3.45 g  $\text{CO}_2$  is produced how many water molecules are also produced?
  - If there are 10.0 g of  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$  and 10.0 g of oxygen reacting. Which is the limiting reagent?
- Use the following equation to answer these questions:  $2 \text{Na}_{(s)} + 2 \text{H}_2\text{O}_{(l)} \rightarrow 2 \text{NaOH}_{(aq)} + \text{H}_{2(g)}$ 
  - How many liters of water is needed to produce 87.69 L  $\text{H}_2$ ?
  - If 90.0 grams of sodium is dropped into 80.0 g of water, how many liters of hydrogen would be produced? (hint this is a limiting reactant problem)
- For the following equation answer these questions:  $3\text{PbCl}_2 + \text{Al}_2(\text{SO}_4)_3 \rightarrow 2 \text{AlCl}_3 + 3\text{PbSO}_4$ 
  - What mass of  $\text{PbCl}_2$  is needed to produce 128 g of  $\text{PbSO}_4$ ?
  - If 25.5 g of  $\text{Al}_2(\text{SO}_4)_3$  reacts with excess  $\text{PbCl}_2$  how much aluminum chloride in grams is produced?
  - What is the limiting reactant if you have 65.45 g of each reactant?
  - What is the excess reactant? How many grams of excess reactant are left over?
  - What is the theoretical yield in grams of the Aluminum Chloride if you react 85 g of  $\text{PbCl}_2$ ?
    - If the actual yield is 42.5 g, what is the % yield?
- Balance the following equation and work the problem.  $\text{FeS} + \text{HCl} \rightarrow \text{H}_2\text{S} + \text{FeCl}_2$ 
  - What mass of Iron(II) Sulfide ( $\text{FeS}$ ) is needed to produce 42.3 L of  $\text{H}_2\text{S}$ ?
  - If 62.43 g of  $\text{HCl}$  is reacted with excess  $\text{FeS}$  how many grams of  $\text{FeCl}_2$  is produced?
  - What is the limiting reactant if you have 23.50 g of each reactant?
  - What is the excess reactant? How many grams of excess reactant are left over?
  - What is the theoretical yield in grams of the  $\text{FeCl}_2$  if you react 22.35 g  $\text{FeS}$  with excess  $\text{HCl}$ ?
    - If the actual yield is 12.5 g, what is the % yield?

# Chemistry Chapter 11 Stoichiometry Study Guide Answers

**Robinson**



## **Chemistry Chapter 11 Stoichiometry Study Guide Answers:**

*The Practice of Chemistry Study Guide & Solutions Manual* Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

**Study Guide/Selected Solutions Manual** Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

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diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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