

Chapter 18

Study Guide - Answers

1. Write the expression for the equilibrium constant for each of the following reactions.



$$K_{\text{eq}} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$



$$K_{\text{eq}} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$



$$K_{\text{eq}} = \frac{[\text{H}^+][\text{CN}^-]}{[\text{HCN}]}$$



$$K_{\text{eq}} = \frac{[\text{H}^+][\text{HSO}_4^-]}{[\text{H}_2\text{SO}_4]}$$



$$K_{\text{eq}} = [\text{Pb}^{2+}][\text{F}^-]^2$$

2. From the data provided below, calculate the value of the equilibrium constant for the reaction.



$$[\text{H}_2] = [\text{Cl}_2] = 1.0 \times 10^{-2}; [\text{HCl}] = 1.0 \times 10^{-4}$$

$$K_{\text{eq}} = \frac{[\text{HCl}]^2}{[\text{H}_2][\text{Cl}_2]}$$

$$K_{\text{eq}} = \frac{[1.0 \times 10^{-4}]^2}{[1.0 \times 10^{-2}][1.0 \times 10^{-2}]} = 0.0001$$



$$[\text{N}_2] = 4.4 \times 10^{-2}; [\text{H}_2] = 1.2 \times 10^{-1}; [\text{NH}_3] = 3.4 \times 10^{-3}$$

$$K_{\text{eq}} = \frac{[\text{NH}_3]^2}{[\text{N}_2][\text{H}_2]^3}$$

$$K_{\text{eq}} = \frac{[3.4 \times 10^{-3}]^2}{[4.4 \times 10^{-2}][1.2 \times 10^{-1}]^3} = 0.152$$

3. For the reaction $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$, $K_{\text{eq}} = 0.2$. At a particular time, the following concentrations are measured: $[\text{N}_2\text{O}_4] = 2.0 \text{ M}$, $[\text{NO}_2] = 0.2 \text{ M}$. Is this reaction at equilibrium? If not, in which direction will the reaction proceed?



$$[\text{N}_2\text{O}_4] = 2.0 \text{ M}$$

$$[\text{NO}_2] = 0.2 \text{ M}$$

$$Q = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]}$$

$$Q = \frac{[0.2]^2}{[2.0]} = 0.02$$

$$K_{\text{eq}} = 0.2$$

$Q < K_{\text{eq}}$, so the reaction will proceed to the right (forward).

Chemistry Chaper 18 Studyguide Answer

H. Stephen Stoker, Danny V. White



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 - The Rise of Digital Reading Chemistry Chapter 18 Studyguide Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 18 Studyguide Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
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 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 18 Studyguide Answer
 - User-Friendly Interface
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 - Personalized Recommendations
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 - Highlighting and Note-Taking Chemistry Chapter 18 Studyguide Answer
 - Interactive Elements Chemistry Chapter 18 Studyguide Answer
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 18 Studyguide Answer
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 - Creating a Diverse Reading Collection Chemistry Chapter 18 Studyguide Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Chemistry Chapter 18 Studyguide Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 18 Studyguide Answer
 - Fact-Checking eBook Content of Chemistry Chapter 18 Studyguide Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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