

AP Chem Chapter 16 Practice Quiz

1. Which statement is FALSE?

- a) If a reaction is thermodynamically spontaneous, it may occur rapidly.
- b) A fast reaction may be thermodynamically spontaneous.
- c) If a reaction is thermodynamically spontaneous, it may occur slowly.
- d) If a reaction is thermodynamically spontaneous, it must have a low activation energy.
- e) Rate of reaction is a kinetic quantity rather than a thermodynamic quantity.

2. Which of the following is a kinetics concept?

- a) free energy
- b) enthalpy
- c) spontaneity
- d) reaction mechanism
- e) entropy

3. A catalyst

- a) is used up in a chemical reaction
- b) changes the value of ΔG of the reaction
- c) is always a solid
- d) does not influence the reaction in any way
- e) changes the activation energy of the reaction

4. What is the entropy change of the reaction below at 298K and 1 atm pressure?



- a) -198.7 J/K
- b) 76.32 J/K
- c) -129.7 J/K
- d) 303.2 J/K
- e) 384.7 J/K

5. If the entropy change for the reaction below at 298 K and 1 atm pressure is 137 J/K, and S=205 J/mol x K for $\text{O}_2(\text{g})$, what is S for $\text{O}_3(\text{g})$?



- a) 364 J/mol x K
- b) 478 J/mol x K
- c) 239 J/mol x K
- d) -117 J/mol x K
- e) -59 J/mol x K

6. The heat of vaporization of methanol, CH_3OH , is 35.20 kJ/mol. Its boiling point is 64.6 degrees Celsius. What is the change in entropy for the vaporization of methanol?

- a) -17.0 J/mol x K
- b) 3.25 J/mol x K
- c) 17.0 J/mol x K
- d) 104 J/mol x K
- e) 543 J/mol x K

7. Which one on the following reactions has a positive entropy change?

- a) $\text{H}_2\text{O}(\text{g}) \longrightarrow \text{H}_2\text{O}(\text{l})$
- b) $\text{BF}_3(\text{g}) + \text{NH}_3(\text{g}) \longrightarrow \text{F}_3\text{BNH}_3(\text{s})$
- c) $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \longrightarrow 2\text{SO}_3(\text{g})$
- d) $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \longrightarrow 2\text{NH}_3(\text{g})$
- e) $2\text{NH}_4\text{NO}_3(\text{s}) \longrightarrow 2\text{N}_2(\text{g}) + 4\text{H}_2\text{O}(\text{g}) + \text{O}_2(\text{g})$

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