

Gas Laws Problem Packet

Gas Laws Packet

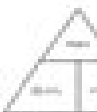

Name _____ Date _____ Period _____

Essential Skills

→ **Standard Pressure**

atm	Standard Pressure
mmHg	760 mm
torr	760
psi	14.7

22.4 L = 1 mol at STP

→ STP (standard temperature and pressure) = 0°C and 1 atm

→ The gas laws:

Boyle's Law	$P_1V_1 = P_2V_2$
Charles' Law	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
Average Gas Law	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
Combined Gas Law	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Combined Gas Law	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Combined Gas Law	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

→ The constant R (used in the ideal gas law) depends on the unit of pressure:

- a. If atm, $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$
- b. If atm, $R = 8.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$

Boyle's Law

- Temperature must always be in Kelvin ($K = ^\circ C + 273$)
- If you are given two different units of pressure, CONVERT one
- Create two columns before solving: what you know, what you're given in the problem. Draw the line that uses the variable you're given, and solve for the unknown. (For Ideal Gas Law, write what you are given. Insert those in equation and solve for unknown)
- Average Gas Law: $\frac{P_1}{T_1} = \frac{P_2}{T_2}$ (Don't forget to use both Kelvin and pressure)
- Combined Gas Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the rate of effusion and molecular masses in the correct units before solving the equation (this is an AVERAGE question). You MUST calculate the square root of volume both sides of the equation before moving things

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A stack of papers with Charles's Law problems, a yellow pencil, and a calculator. The top paper shows the formula $\frac{V_1}{T_1} = \frac{V_2}{T_2}$ and a problem: "A gas sample with a volume of 1.0 L at 273 K is heated to 546 K. What is the new volume?" The second paper shows the same formula and a problem: "A gas sample with a volume of 2.0 L at 273 K is cooled to 136 K. What is the new volume?" The third paper shows the same formula and a problem: "A gas sample with a volume of 3.0 L at 273 K is cooled to 136 K. What is the new volume?" A yellow pencil is resting on the top paper, and a calculator is in the foreground.

All Gas Laws Included!

Chemistry B Gases Packet Answer Key

SB Merriam



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