

Chemistry B Final Study Guide

8

- What is a precipitation reaction? Provide an example.
- Summarize the simple solubility rules for ionic compounds.
- What is a combustion reaction? Write an equation that illustrates a combustion reaction.
- Give an example of a synthesis reaction and of a decomposition reaction.
- Classify the reaction represented by each of the following chemical equations in as many ways as possible based sections 8.6 and 8.7. Balance each equation.

- $\text{NaOH(s)} + \text{CuSO}_4\text{(aq)} \rightarrow \text{Cu(OH)}_2\text{(s)} + \text{Na}_2\text{SO}_4\text{(aq)}$
- $\text{HI(aq)} + \text{KOH(aq)} \rightarrow \text{KI(aq)} + \text{H}_2\text{O(l)}$
- $\text{FeO(s)} + \text{HNO}_3\text{(aq)} \rightarrow \text{Fe(NO}_3)_2\text{(aq)} + \text{H}_2\text{O(l)}$
- $\text{C}_{12}\text{H}_{22}\text{O}_{11}\text{(s)} \rightarrow \text{C(s)} + \text{H}_2\text{O(g)}$
- $\text{B(s)} + \text{O}_2\text{(g)} \rightarrow \text{B}_2\text{O}_3\text{(s)}$

-Know the type of reactions listed in sections 8.6 and 8.7.

9

- Considering the reaction represented by the (unbalanced) equation



Determine the number of moles of $\text{NH}_3\text{(g)}$ that can be produced from the following:

- 0.20 mol $\text{N}_2\text{(g)}$ reacts completely with $\text{H}_2\text{(g)}$.
- 0.30 mol $\text{H}_2\text{(g)}$ reacts completely with $\text{N}_2\text{(g)}$.

- Considering the reaction represented by the (unbalanced) equation



Determine the mass of $\text{H}_2\text{(g)}$ that can be produced from the following:

- 10.0 g Mg(s) reacts completely with HCl(aq) .
- 20.0 g HCl(aq) reacts completely with Mg(s) .

10

- How is the concept of energy defined?
- What does temperature measure?
- Convert the temperatures to Kelvin:
 - 25°C
 - 56°C
 - 96°C
- Explain what is meant by the terms exothermic and endothermic.
- What is meant by the specific heat capacity of a material

-Calculate the mass (in grams) of each of the following substances that could be warmed over the indicated temperature range by application of exactly 1.0 kJ of energy.

- water, from 15°C to 42°C $s = 4.184 \text{ J/g } ^\circ\text{C}$
- iron, from 25°C to 125°C $s = 0.45 \text{ J/g } ^\circ\text{C}$
- carbon, from -10°C to 47°C $s = 0.71 \text{ J/g } ^\circ\text{C}$

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