

- ____ 9. At 700 °C, the equilibrium constant (K_c) for the reaction $\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \rightleftharpoons 2 \text{NH}_3(\text{g})$ is 2.37×10^{-3} . If 2.4 mol of N_2 , 3.20 mol of H_2 and 0.632 mole of NH_3 are mixed in a 2.00 L container at 700 °C:
a. the value of Q is unknown. b. the value of Q is greater than K_c . c. the value of Q is equal to K_c .
d. the reaction is at equilibrium. e. the value of Q is less than K_c .
- ____ 10. The only factor that causes a change in the numerical value of an equilibrium constant is a(n):
a. change in volume. b. change in pressure. c. change in concentrations. d. addition of a catalyst.
e. change in temperature.
- ____ 11. Which of the items listed below would increase the yield of NO_2 in the following equilibrium reaction?
- $$2 \text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2 \text{NO}_2(\text{g}) \quad (\text{exothermic})$$
- I. Add more O_2 to the equilibrium mixture.
II. Warm the reaction mixture.
III. Increase the pressure by decreasing the volume of the reaction vessel.
- a. II only b. I and II c. I and III d. II and III e. All of these.
- ____ 12. For the reaction $\text{H}_2(\text{g}) + \text{CO}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{g}) + \text{CO}(\text{g})$ at 700 °C, $K_c = 0.534$. Calculate the concentration of H_2 at equilibrium if a mixture of 0.200 mol of H_2O , 0.200 mol CO , 0.750 mol H_2 and 0.750 mol CO_2 is heated to 700 °C in a 1.00 liter container.
a. 0.500 M b. 0.100 M c. 0.573 M d. 0.132 M e. 0.735 M
- ____ 13. The conjugate acid of HPO_4^{2-} is:
a. H_3PO_4 b. H_2PO_4^- c. PO_4^{3-} d. H^+ e. none of these
- ____ 14. If the reaction below proceeds, which of the following is true?
- $$\text{NH}_3 + \text{C}_2\text{H}_5\text{OH} \rightarrow \text{NH}_2^- + \text{C}_2\text{H}_5\text{OH}_2^+$$
- a. $\text{C}_2\text{H}_5\text{OH}$ is an acid. b. NH_2^- is the conjugate base of NH_3 . c. NH_3 is a base. d. $\text{C}_2\text{H}_5\text{OH}_2^+$ is a base.
e. none of these
- ____ 15. Which aqueous solution(s) below is(are) considered basic?
- I. A solution with a pH = 6
II. A solution with $[\text{OH}^-] = 1 \times 10^{-7}$
III. A solution with $[\text{H}_3\text{O}^+] = 1 \times 10^{-9}$
- a. I only b. II only c. III only d. I and II e. I and III
- ____ 16. The pH of a diet Coca Cola solution is 2.72. What is the concentration of $[\text{OH}^-]$ present in this solution?
a. $[\text{OH}^-] = 5.25 \times 10^{-12} \text{ M}$ b. $[\text{OH}^-] = 1.91 \times 10^{-2} \text{ M}$ c. $[\text{OH}^-] = 1.3 \times 10^{-12} \text{ M}$ d. $[\text{OH}^-] = 7.6 \times 10^{-4} \text{ M}$
e. $[\text{OH}^-] = 3.16 \times 10^{10} \text{ M}$

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F Rizvi



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