

Explain:

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1. **Explain the following observations:**
 (a) **Reaction of iron with steam:**

$$3\text{Fe} + 4\text{H}_2\text{O} \xrightarrow{\text{steam}} \text{Fe}_3\text{O}_4 + 4\text{H}_2$$

 (b) **Reaction of iron with dilute acids:**

$$\text{Fe} + 2\text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2$$

 (c) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

2. **Explain the following observations:**
 (a) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (b) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (c) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

3. **Explain the following observations:**
 (a) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (b) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (c) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

4. **Explain the following observations:**
 (a) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (b) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (c) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

5. **Explain the following observations:**
 (a) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (b) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (c) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

6. **Explain the following observations:**
 (a) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (b) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (c) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

7. **Explain the following observations:**
 (a) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (b) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (c) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

8. **Explain the following observations:**
 (a) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (b) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (c) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

9. **Explain the following observations:**
 (a) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (b) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (c) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

10. **Explain the following observations:**
 (a) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (b) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

 (c) **Reaction of iron with concentrated acids:**

$$\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$$

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Richard O. Straub



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