

2. Write the formulas for the following IONIC compounds.

a) Potassium nitride  $K_3N$

f) Zinc oxalate  $ZnC_2O_4$

b) Iron(III) sulfite  $Fe_2(SO_3)_3$

g) Magnesium phosphate  $Mg_3(PO_4)_2$

c) Calcium oxide  $CaO$

h) Cobalt(III) permanganate  $Co(MnO_4)_3$

d) Tin(IV) carbonate  $Sn(CO_3)_2$

i) Silver sulfate  $Ag_2SO_4$

e) Gallium chlorate  $Ga(ClO_3)_3$

j) Lead(IV) fluoride  $PbF_4$

3. Name the following IONIC compounds.

a)  $Na(BrO_3)$  sodium bromate

k)  $SnF_2$  tin(II) fluoride

b)  $K_3(AsO_4)$  potassium arsenate

l)  $Hg_2Br_2$  mercury(I) bromide

c)  $Sr(NO_2)_2$  strontium nitrite

m)  $Co(PO_4)$  cobalt(III) phosphate

n)  $Hg_2Cl_2$  mercury(I) chloride

o)  $PbCl_2$  lead(II) chloride

p)  $FeCl_3$  iron(III) chloride

q)  $FeCl_2$  iron(II) chloride

r)  $Na_3P$  sodium phosphide

s)  $CaCl_2$  calcium chloride

t)  $Al_2O_3$  aluminum oxide

u)  $CrCl_3$  chromium(III) chloride

v)  $AgCl$  silver chloride

w)  $SnCl_4$  tin(IV) chloride

x)  $SnCl_2$  tin(II) chloride

y)  $CoCl_2$  cobalt(II) chloride

z)  $AlCl_3$  aluminum chloride

aa)  $AgNO_3$  silver nitrate

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