

# 7

## Chemical Equilibrium

1. For the reversible reaction,

$$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g}) + \text{heat}$$
the equilibrium shifts in forward direction  
[2014]

- (a) by increasing the concentration of  $\text{NH}_3(\text{g})$   
(b) by decreasing the pressure  
(c) by decreasing the concentrations of  $\text{N}_2(\text{g})$  and  $\text{H}_2(\text{g})$   
(d) by increasing pressure and decreasing temperature
2. Using the Gibbs energy change,  $\Delta G^\circ = +63.3 \text{ kJ}$  for the following reaction,  
$$\text{Ag}_2\text{CO}_3(\text{s}) \rightleftharpoons 2\text{Ag}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq})$$
the  $K_{\text{sp}}$  of  $\text{Ag}_2\text{CO}_3(\text{s})$  in water at  $25^\circ\text{C}$  is  
( $R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$ ) [2014]
- (a)  $3.2 \times 10^{-26}$   
(b)  $8.0 \times 10^{-12}$   
(c)  $2.9 \times 10^{-3}$   
(d)  $7.9 \times 10^{-7}$

3.  $\text{KMnO}_4$  can be prepared from  $\text{K}_2\text{MnO}_4$  as per reaction,



The reaction can go to completion by removing  $\text{OH}^-$  ions by adding [2013]

- (a)  $\text{HCl}$   
(b)  $\text{KOH}$   
(c)  $\text{CO}_2$   
(d)  $\text{SO}_2$
4. The value of  $\Delta H$  for the reaction,  
$$\text{X}_2(\text{g}) + 4\text{Y}_2(\text{g}) \rightleftharpoons 2\text{XY}_4(\text{g})$$
is less than zero. Formation of  $\text{XY}_4(\text{g})$  will be favoured at [2011]
- (a) low pressure and low temperature  
(b) high temperature and low pressure  
(c) high pressure and low temperature  
(d) high temperature and high pressure
5. For the reaction,  
$$\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$$
, the equilibrium constant is  $K_1$ . The equilibrium constant is  $K_2$  for the reaction,  
$$2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$$
. What is  $K$  for the reaction,  $\text{NO}_2(\text{g}) \rightleftharpoons \frac{1}{2}\text{N}_2(\text{g}) + \text{O}_2(\text{g})$ ? [2011]
- (a)  $1/(4K_1K_2)$  (b)  $[1/(K_1K_2)]^{1/2}$   
(c)  $1/(K_1K_2)$  (d)  $1/8K_1K_2$
6. In which of the following equilibrium  $K_c$  and  $K_p$  are not equal? [2010]
- (a)  $2\text{NO}(\text{g}) \rightleftharpoons \text{N}_2(\text{g}) + \text{O}_2(\text{g})$   
(b)  $\text{SO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons \text{SO}_3(\text{g}) + \text{NO}(\text{g})$   
(c)  $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$   
(d)  $2\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{CO}_2(\text{g})$
7. The dissociation constants for acetic acid and  $\text{HCN}$  at  $25^\circ\text{C}$  are  $1.5 \times 10^{-5}$  and  $4.5 \times 10^{-10}$ , respectively. The equilibrium constant for the equilibrium,  
$$\text{CN}^- + \text{CH}_3\text{COOH} \rightleftharpoons \text{HCN} + \text{CH}_3\text{COO}^-$$
would be [2009]
- (a)  $3.0 \times 10^5$  (b)  $3.0 \times 10^{-5}$   
(c)  $3.0 \times 10^{-4}$  (d)  $3.0 \times 10^4$

# Chemical Equilibrium Solutions Manual Answers

**C Cleary**



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