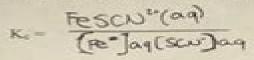
#### Report Sheet - Experiment 13-Chemical Equilibrium

- 1. Write the chemical equation for the reversible reaction studied in this experiment:

  Fe \*(aq)+ SCN \* > Fe (SCN) \*(aq)
- 2. Write the equilibrium expression for the reaction studied in this experiment:



3. Complete the following table from your observations of test tubes #1-#6:

Test Tube	Change	Initial Color	Final Color (darker/lighter)
	Initial color before anything is added		
1	Added Fe(NO <sub>3</sub> ) <sub>3</sub>	Reddish color	dorker
2	Added KSCN	Reddish color	darker
3	Added NH <sub>4</sub> CI	Reddish color	Lighter
4	Added AgNO	Reddish color	Lighter
5	Cold	Reddish color	darker
6	Hot	Reddish color	lighter

4. Complete the following table for test tubes #2 - 4:

Test Tube	What concentration was changed and how it was changed?	The equilibrium shifted to left/right/no change	What evidence for the direction of shift do you have?
2	Fesch t is a mix of an iron I relaish color + thiconate has no color	Shifted Paykt	principle but react on from lett to right accus was when more reactant add
3	(SCN) was	Shifted Right	the product goes up
4	Fezt concentration was	Shifetell Just+	Fe COL WILL Form which leads to a lower Fe to concentration

# **Chemical Equilibrium Lab Report Answers**

**Brendan G. Carr** 

## **Chemical Equilibrium Lab Report Answers:**

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#### **Table of Contents Chemical Equilibrium Lab Report Answers**

- 1. Understanding the eBook Chemical Equilibrium Lab Report Answers
  - The Rise of Digital Reading Chemical Equilibrium Lab Report Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemical Equilibrium Lab Report Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Equilibrium Lab Report Answers
  - User-Friendly Interface

- 4. Exploring eBook Recommendations from Chemical Equilibrium Lab Report Answers
  - Personalized Recommendations
  - Chemical Equilibrium Lab Report Answers User Reviews and Ratings
  - Chemical Equilibrium Lab Report Answers and Bestseller Lists
- 5. Accessing Chemical Equilibrium Lab Report Answers Free and Paid eBooks
  - Chemical Equilibrium Lab Report Answers Public Domain eBooks
  - Chemical Equilibrium Lab Report Answers eBook Subscription Services
  - Chemical Equilibrium Lab Report Answers Budget-Friendly Options
- 6. Navigating Chemical Equilibrium Lab Report Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Chemical Equilibrium Lab Report Answers Compatibility with Devices
  - Chemical Equilibrium Lab Report Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Equilibrium Lab Report Answers
  - Highlighting and Note-Taking Chemical Equilibrium Lab Report Answers
  - Interactive Elements Chemical Equilibrium Lab Report Answers
- 8. Staying Engaged with Chemical Equilibrium Lab Report Answers
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Equilibrium Lab Report Answers
- 9. Balancing eBooks and Physical Books Chemical Equilibrium Lab Report Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Equilibrium Lab Report Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Equilibrium Lab Report Answers
  - Setting Reading Goals Chemical Equilibrium Lab Report Answers
  - Carving Out Dedicated Reading Time

- 12. Sourcing Reliable Information of Chemical Equilibrium Lab Report Answers
  - Fact-Checking eBook Content of Chemical Equilibrium Lab Report Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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