



# CHEMISTRY

## CHEAT SHEET

# EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol  $\rightleftharpoons$ .

<p><b>What is Equilibrium?</b></p> <p>Chemical Equilibrium is a state where the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol <math>\rightleftharpoons</math>.</p> <p><b>Equilibrium Constant (<math>K_c</math>)</b></p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant <math>K_c</math> is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p><b>Equilibrium Constant (<math>K_c</math>)</b></p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant <math>K_c</math> is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$									
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<p><b>Le Chatelier's Principle</b></p> <p>When a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system shifts in the direction that tends to counteract the change and restore equilibrium.</p> <p><b>Equilibrium Constant (<math>K_c</math>)</b></p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant <math>K_c</math> is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p><b>Le Chatelier's Principle</b></p> <table><tr><th>Change in Conditions</th><th>Effect on Equilibrium</th></tr><tr><td>Concentration</td><td>Increasing concentration of a reactant shifts the equilibrium to the right. Increasing concentration of a product shifts the equilibrium to the left.</td></tr><tr><td>Pressure</td><td>Increasing pressure shifts the equilibrium to the side with fewer moles of gas. Decreasing pressure shifts the equilibrium to the side with more moles of gas.</td></tr><tr><td>Temperature</td><td>Increasing temperature shifts the equilibrium to the endothermic side. Decreasing temperature shifts the equilibrium to the exothermic side.</td></tr></table>		Change in Conditions	Effect on Equilibrium	Concentration	Increasing concentration of a reactant shifts the equilibrium to the right. Increasing concentration of a product shifts the equilibrium to the left.	Pressure	Increasing pressure shifts the equilibrium to the side with fewer moles of gas. Decreasing pressure shifts the equilibrium to the side with more moles of gas.	Temperature	Increasing temperature shifts the equilibrium to the endothermic side. Decreasing temperature shifts the equilibrium to the exothermic side.
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Le Chatelier's Principle: the effect of changing conditions on equilibrium										
Condition	Effect on Equilibrium	Example								
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Pressure	<ul style="list-style-type: none"><li>Increasing pressure shifts the equilibrium to the side with fewer moles of gas.</li><li>Decreasing pressure shifts the equilibrium to the side with more moles of gas.</li></ul>	For the reaction: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing the pressure shifts the equilibrium to the right, producing more $NH_3$ .								
Temperature	<ul style="list-style-type: none"><li>Increasing temperature shifts the equilibrium to the endothermic side.</li><li>Decreasing temperature shifts the equilibrium to the exothermic side.</li></ul>	For the reaction: $N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing the temperature shifts the equilibrium to the left, producing less $NH_3$ .								

# Chemical Equilibrium Study Guide Answer Sheet

**Yijin Wang**



## Chemical Equilibrium Study Guide Answer Sheet:

*Chemical Principles Study Guide/Solutions Manual* John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

**Chemical Principles Student's Study Guide & Solutions Manual** John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

**CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE** Kevin Smith, 2024-02-01 A comprehensive summary of Grade 11 12 Physics Simple logical summaries with example exam questions and work through solutions The book covers the fundamentals of Grade 11 12 Physics and complements the material in any class text

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Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement *The Complete*

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