

517. At a temperature of 1,800 K, the reactions represented by the following two equations predominate in air at 1.0 atm pressure:



Assume the following:

For air:

$$\text{O}_2 = 21 \text{ vol\%}$$

$$\text{N}_2 = 79 \text{ vol\%}$$

x = NO formed at equilibrium, mol

y = NO₂ formed at equilibrium, mol

Which of the following expressions for the equilibrium constant is correct for Reaction 1?

(A)
$$\frac{x}{(0.79 - x/2 - y/2)^{1/2} (0.21 - x/2 - y)^{1/2}}$$

(B)
$$\frac{x}{(1 - y/2)(0.79 - x/2 - y/2)^{1/2} (0.21 - x/2 - y)^{1/2}}$$

(C)
$$\frac{x(1 - y/2)}{(0.79 - x/2 - y/2)^{1/2} (0.21 - x/2 - y)^{1/2}}$$

(D)
$$\frac{x}{(0.79 - x/2 - y)^{1/2} (0.21 - x/2 - y/2)^{1/2}}$$

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