

Name key

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. When an atom loses an electron it becomes a positive ion (cation), and when it gains an electron it becomes a negative ion (anion).
2. An element in the halogen group has 7 valence electrons, and wants to gain 1 more to have 8. This is because of the octet rule.
3. An ionic bond is the attraction between oppositely charged ions, AKA a metal and a nonmetal.
4. In a covalent bond, atoms share electrons, whereas in an ionic bond, atoms transfer electrons.
5. (Honors only) When electrons are shared unequally in a covalent bond it is polar, and when they're shared equally it is nonpolar. You can predict if a compound is ionic, polar covalent or nonpolar covalent by subtracting the electronegativities.
6. An atom can have up to 8 valence electrons. The family (column) on the periodic table that all have 8 valence electrons is called noble gases.
7. Atoms can gain, lose, or share valence electrons to become more stable.

B. Formulas for Ionic compounds - Fill in the chart below

	Magnesium	Fluorine	Nitrogen	Lithium
# OF VALENCE ELECTRONS	<u>2</u>	<u>7</u>	<u>5</u>	<u>1</u>
OXIDATION #	<u>+2</u>	<u>-1</u>	<u>-3</u>	<u>+1</u>

2. What is the **chemical formula** for each of the following ionic compounds?
(*hint: you can either draw the Lewis structures and use arrows to show the transfer of electrons, OR use the periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.*)

- a) Potassium and Chlorine - KCl
- b) Calcium and Bromine - CaBr₂
- c) Magnesium and oxygen - MgO

Chemical Bonding Study Guide With Answers

B Lingard



Chemical Bonding Study Guide With Answers:

Hazmat Chemistry Study Guide (Second Edition) Jill Meryl Levy, 2011 Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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Table of Contents Chemical Bonding Study Guide With Answers

1. Understanding the eBook Chemical Bonding Study Guide With Answers
 - The Rise of Digital Reading Chemical Bonding Study Guide With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Bonding Study Guide With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemical Bonding Study Guide With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Bonding Study Guide With Answers
 - Personalized Recommendations
 - Chemical Bonding Study Guide With Answers User Reviews and Ratings
 - Chemical Bonding Study Guide With Answers and Bestseller Lists
5. Accessing Chemical Bonding Study Guide With Answers Free and Paid eBooks
 - Chemical Bonding Study Guide With Answers Public Domain eBooks
 - Chemical Bonding Study Guide With Answers eBook Subscription Services
 - Chemical Bonding Study Guide With Answers Budget-Friendly Options
6. Navigating Chemical Bonding Study Guide With Answers eBook Formats

- ePub, PDF, MOBI, and More
- Chemical Bonding Study Guide With Answers Compatibility with Devices
- Chemical Bonding Study Guide With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Bonding Study Guide With Answers
 - Highlighting and Note-Taking Chemical Bonding Study Guide With Answers
 - Interactive Elements Chemical Bonding Study Guide With Answers
- 8. Staying Engaged with Chemical Bonding Study Guide With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Bonding Study Guide With Answers
- 9. Balancing eBooks and Physical Books Chemical Bonding Study Guide With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Bonding Study Guide With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Bonding Study Guide With Answers
 - Setting Reading Goals Chemical Bonding Study Guide With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Bonding Study Guide With Answers
 - Fact-Checking eBook Content of Chemical Bonding Study Guide With Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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