

Chemical Reactions and Stoichiometry

1. Given the equation $3A + B \rightarrow C + D$, you react 1 mole of A with 3 moles of B. True or false: B is the limiting reactant because you have fewer moles of B than A.

Answer: False

$$1 \text{ mole A} \frac{1 \text{ mole B}}{3 \text{ mole A}} = 0.3 \text{ moles B needed. Have 3 moles B so A is limiting}$$

2. NaHCO_3 is the active ingredient in baking soda. How many grams of oxygen are in 0.35 g of NaHCO_3 ?

- a) 0.012 g
- b) 0.022 g
- c) 0.067 g
- d) 0.20 g

ANS: d)

$$0.35 \text{ g NaHCO}_3 \left(\frac{1 \text{ mole NaHCO}_3}{84.07 \text{ g NaHCO}_3} \right) \left(\frac{3 \text{ moles O}}{1 \text{ mole NaHCO}_3} \right) \left(\frac{16.00 \text{ g}}{1 \text{ mole O}} \right) = 0.199 = \boxed{0.20 \text{ g}}$$

3. Suppose the reaction $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$ is carried out starting with 103 g of $\text{Ca}_3(\text{PO}_4)_2$ and 75.0 g of H_2SO_4 . How much phosphoric acid will be produced?

- a) 74.9 g
- b) 50.0 g
- c) 112 g
- d) 32.5 g

ANS: b)

$$\begin{aligned} 103 \text{ g Ca}_3(\text{PO}_4)_2 & \frac{1 \text{ mole}}{310.1875 \text{ g Ca}_3(\text{PO}_4)_2} = 0.332 \text{ moles Ca}_3(\text{PO}_4)_2 \\ 75.0 \text{ g H}_2\text{SO}_4 & \frac{1 \text{ mole H}_2\text{SO}_4}{98.082 \text{ g H}_2\text{SO}_4} = 0.765 \text{ moles H}_2\text{SO}_4 \text{ L.R.} \\ \text{Ca}_3(\text{PO}_4)_2 \text{ needed: } & 0.765 \text{ moles H}_2\text{SO}_4 \frac{1 \text{ mole Ca}_3(\text{PO}_4)_2}{3 \text{ moles H}_2\text{SO}_4} = 0.255 \text{ moles Have excess} \\ 0.765 \text{ moles H}_2\text{SO}_4 & \frac{2 \text{ moles H}_3\text{PO}_4}{3 \text{ moles H}_2\text{SO}_4} \frac{97.994 \text{ g H}_3\text{PO}_4}{1 \text{ mole}} = 49.97 \text{ g} = \boxed{50.0 \text{ g}} \end{aligned}$$

Chemfiesta Stoichiometry Lab Answers

Gregory Gebhart



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A Stoichiometry Unit David Callaghan, 2004 **Stoichiometry Unit Project** Luann Marie Decker, 1998 *Chemistry Study Guide* Gregory Gebhart, 2017-09-21 Solutions to stoichiometry and conservation of mass problems Laboratory exercises

You Can Do Chemistry Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

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