

For each problem below, list each variable given and which one is unknown, choose the correct equation and plug in the numbers to solve for the unknown.

2. 234 kPa of Argon is at a temp of 255 K. If the temp changes to 100 K, what is the new pressure?

$$P_1 = 234 \text{ kPa}$$

$$T_1 = 255 \text{ K}$$

$$T_2 = 100 \text{ K}$$

$$P_2 = ?$$

$$\boxed{100 \text{ K} \left(\frac{234 \text{ kPa}}{255 \text{ K}} \right) = \frac{P_2}{100 \text{ K}} (100 \text{ K})}$$

$$P_2 = 91.765 \text{ kPa}$$

3. A 6.5 liter tank of methane gas has compressed gas in it so the pressure is 2500 kPa. What will the pressure of the gas be if the gas is put into a 3.25 liter tank?

$$V_1 = 6.5 \text{ L}$$

$$P_1 = 2500 \text{ kPa}$$

$$V_2 = 3.25 \text{ L}$$

$$P_2 = ?$$

$$\boxed{\frac{(2500)(6.5)}{3.25} = \frac{P_2 (3.25)}{3.25}}$$

$$P_2 = 5000 \text{ kPa}$$

4. Neon gas at a pressure of 25 kPa occupies 444 liters at 333 K. What is the pressure of the gas at 200 K in a volume of 30 liters?

Chemfiesta Gas Law Practice Boyles Law

Patrick H. Martin, Bruce M. Kramer



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