PREDICTING PRODUCTS REVIEW SHEET

The following reactants already have balancing coefficients for their likely products. Write the balanced products with states or "NR" if no reaction will take place. **Underneath** each reaction tell which type of reaction is taking place. There may be more than one.

- Mg(OH)_{2 (MQ)} + 2 HNO_{3 (MQ)} → 2 H₂O (I) + Mg(NO₃)_{2 (MQ)}
 This is Double Replacement. More specifically, this is the special type called and Acid/Base reaction to form water. It is also called Neutralization.
- 3 KOH (aq) + H₃PO₄ (aq) → K₃PO₄ (aq) + 3 H₂O (t)
 Double Replacement Acid/Base Reaction
- 3. $2 C_8 H_{18 (g)} + 25 O_{2 (g) [excess]} \rightarrow 16 CO_{2 (g)} + 18 H_2O_{(g)}$ Combustion
- 4. $Sn(C_2H_3O_2)_{2 \text{ (sq)}} + H_2S_{\text{ (sq)}} \rightarrow 2 HC_2H_3O_{2 \text{ (sq)}} + SnS_{\text{ (s)}}$ Double Replacement – You may not know it but SnS is a solid so this reaction will occur.
- 5. 2 NaI $_{(96)}$ Br $_{2\,(9)}$ \rightarrow 2 NaBr $_{(96)}$ + I $_{2\,(8)}$ Single Replacement Bromine is more active than iodine.
- 2 Zn (s) + O_{2 (g)} → 2 ZnO (s)
 Synthesis − Type #1
- 7. Na₂O (aq) + H₂O (I) → _____
- 8. $Ca(OH)_{2(aq)} + 2 HCI_{(aq)} \rightarrow CaCI_{2(aq)} + 2 H_2O_{(i)}$ Double Replacement – Acid/Base
- 9. 2 NaCl (s) →elec→ _____
- 10. 2 AlBr_{3 (aq)} + 3 Cl_{2 (g)} → 2 AlCl_{3 (aq)} + 3 Br_{2 (f)} Single Replacement
- 11. 2 KClO_{3 (s)} →Δ→ _____

Chemfiesta Predicting Reaction Products Answer Key

LL Leslie

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