

Data/Observations

In the table below, write your observations of the reactants before and after the reaction occurs.

Reaction	Observations Before Reaction	Observations After Reaction
A	Copper pellets have a shiny, orange appearance. Silver nitrate is a cloudy grey liquid.	Copper pellets become coated with a grey substance, causing them to lose their color.
B	Zinc is a hard, shiny grey metal. Hydrochloric acid is a clear liquid.	The zinc begins to bubble and fizz as the zinc metal slowly begins to disappear.
C	Lead nitrate and potassium iodide are both clear liquids before being mixed.	When the two chemicals are mixed, they react to form a bright yellow substance.
D	Copper (II) sulfate is a bright blue substance that is shaped like small stones or crystals.	The copper (II) sulfate changes from a blue to white color when heated in the Bunsen burner.
E	Magnesium is a grey, shiny metal that is cut into strips.	When heated in the flame, the magnesium releases a bright white light and turns to dust.

The table below lists the equations for the five chemical reactions that were completed in the lab. Identify the type of reaction completed in each step of the lab.

Reaction	Equation	Type of Reaction
A	$\text{Cu} + 2 \text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2 \text{Ag}$	Single Displacement
B	$\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$	Single Displacement
C	$\text{Pb}(\text{NO}_3)_2 + 2\text{KI} \rightarrow \text{PbI}_2 + 2\text{KNO}_3$	Double Displacement
D	$\text{CuSO}_4 \rightarrow \text{CuO} + \text{SO}_2$	Decomposition
E	$2 \text{Mg} + \text{O}_2 \rightarrow 2 \text{MgO}$	Synthesis

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