

ChemActivity 5: Resonance

PART A: DRAWING RESONANCE STRUCTURES

(Why is baking soda $[\text{NaHCO}_3]$ a weak base even though it has an O with a -1 formal charge?)

Model 1: +1 Counterions

We have encountered many -1 anions in this course: $\text{HO}^\ominus$, $\text{Cl}^\ominus$, $(\text{CH}_3)_3\text{N}^\ominus$, $\text{HCO}_2^\ominus$, etc.

If you go to the stockroom you will not find a bottle of " $\text{HO}^\ominus$ " (hydroxide) on the shelf because every anion must exist in close proximity to a counterion, an ion that balances out the charge.

In this book we will generally assume that positive counterions are inert and that their only function is to balance the negative charge of the anion.

Also assume that all ions appearing in this book have a counterion, even if one is not specified.

If ions in a sample did not have counterions, electrical charge would build up, and the sample would find a way to "ground" itself. For example: ion flow across a salt bridge in a battery, or lightning—the result of ion imbalance in a cloud.

Critical Thinking Questions

1. Which elements on the periodic table (other than H) are likely to form a +1 cation?

Li, Na, K, Rb, Cs, Fr

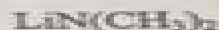
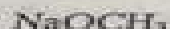
2. You will not find "hydroxide" in the stockroom, but you will find sodium hydroxide (NaOH) and potassium hydroxide (KOH). Lithium hydroxide (LiOH) is expensive and used in spacecraft air filters since hydroxide reacts with carbon dioxide, and lithium is lighter than sodium or potassium. Cesium and francium hydroxides are very expensive and little used. Is this information consistent with your answer to the previous question?

Information: $\text{Li}^\oplus$, $\text{Na}^\oplus$, $\text{K}^\oplus$ (= ions) vs. Li° , Na° , K° (= metals)

The use of a superscript zero after an element (e.g., Li°) indicates the element is in an uncharged metallic state (the zero oxidation state). Li° , Na° , and K° are all highly reactive metals we will occasionally encounter in this course—look for the zero superscript ($^\circ$).

If you don't see the zero superscript ($^\circ$) after Li, Na, or K, assume it is an inert +1 counterion!

3. For each condensed formula, draw a line-bond structure of the anion, and next to it draw the +1 counterion as a separate ion.



Chemactivity 5 Answers

**Norbert J. Pienta, Melanie M.
Cooper, Thomas J. Greenbowe**



Chemactivity 5 Answers:

Chemactivity 5 Answers Book Review: Unveiling the Power of Words

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