

CHEM 108
ChemActivity 23: Intermolecular Forces
Solution Key

The Model: Dipole-Dipole Forces

Within a polar molecule there is a separation of charge. Neighboring molecules in a bulk sample of a polar compound will arrange themselves in order to maximize attractive forces and to minimize repulsive forces between the molecules. (In a liquid there is still enough thermal energy that the molecules are constantly moving past each other; so this is a "time-averaged" picture we're considering.) The net attractive forces between polar molecules are called **dipole-dipole forces**, which are typically only about 1% as strong (~ 4 kJ/mol) as a covalent bond.



Key Questions

1. When a substance boils, at the surface of the liquid inter-particle attractive forces must be broken as particles go from the liquid into the corresponding vapor phase.



Coulomb's Law
$$F = k \frac{Q_+ Q_-}{r^2}$$



- a. The attractive forces between pairs of particles in a polar molecular substance (on the left, in which $\delta+$ represents a partial positive charge and $\delta-$ represents a partial negative charge) is weaker stronger than that in an ionic substance (on the right). (Circle your answer.)
- b. As attractive forces between neighboring molecules in a bulk sample get stronger, it will require less more heat energy to separate molecules from each other. (Circle your answer.)
- c. According to Coulomb's Law, the boiling point of molecular substances should be lower higher than the boiling point of ionic substances. (Circle your answer. Is this consistent with observation?)

Chemactivity 23 The Dipole Moment Answers

Lauren Gardner



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