

# CHEM 101 Worksheet, Chapter 3: Mass Relationships in Chemical Reactions

Multiple choice questions: circle the correct answer(A-E)

Show your work questions (p. 2): show all work! Use the amu for each element from the periodic table!

$N_A = 6.02 \times 10^{23}$

1. Atom X has two isotopes: 25.0 wt% is  $^{37}\text{X}$  and 75.0% is  $^{36}\text{X}$ . What is the average molar mass of X?

- A. 36.0 amu      B. 72.0 amu      C. 36.5 amu      D. 35.5 amu      E. 34.5 amu

2. The molar mass (in amu) of  $\text{ZnSO}_4 \cdot 6\text{H}_2\text{O}$  =

- A. 161.5      B. 179.5      C. 269.5      D. 53.5      E. 215.5

3. How many atoms of Na in 5.10 moles of  $\text{Na}_2\text{CO}_3$ ?

- A.  $3.07 \times 10^{23}$       B.  $6.14 \times 10^{24}$       C.  $6.14 \times 10^{23}$       D.  $1.84 \times 10^{25}$       E. 10.2

4. What is the mass in grams of  $1.00 \times 10^{12}$  atoms of lead, Pb?

- A.  $3.44 \times 10^{-10}\text{g}$       B.  $2.07 \times 10^{114}\text{g}$       C.  $1.25 \times 10^{38}\text{g}$       D.  $3.44 \times 10^{-34}\text{g}$       E.  $1.25 \times 10^{14}\text{g}$

5. 0.2 mol  $\text{Na}_2\text{CO}_3$  = 4.24 g  $\text{Na}_2\text{CO}_3$  = 0.051 mol

- A. 530g, 449 mol      B. 21.2g, 449 mol      C. 530g, 0.051 mol      D. 21.2g, 0.040 mol      E. 16.6g, 0.051 mol

6. The wt% composition in tin(IV)oxide is 88% Sn 12% O (Sn = tin, element # 50 )

- A. 21%Sn, 79%O      B. 88%Sn, 12%O      C. 12%Sn, 88%O      D. 94%Sn, 6%O      E. 79% Sn, 21%O

7. A compound has the following wt% composition: 2.1% H, 65.3% O, 32.6% S. What is its empirical formula?

- A.  $\text{H}_2\text{SO}_4$       B.  $\text{H}_4\text{S}_2\text{O}_8$       C.  $\text{S}_4\text{HO}_2$       D.  $\text{H}_2\text{SO}_2$       E.  $\text{HO}_4\text{S}$

8. How many gram Na in 5.00g  $\text{Na}_2\text{SO}_4$ ?

- A. 0.035 g      B. 1.31 g      C. 1.62 g      D. 0.97 g      E. 0.81 g

9. A compound containing only carbon and hydrogen has 85.7 wt% C and 14.3 wt% H. Its molar mass is 56 amu. What is its molecular formula?

- A.  $\text{C}_4\text{H}_8$       B.  $\text{C}_2\text{H}_4$       C.  $\text{C}_3\text{H}$       D.  $\text{CH}_2$       E. CH

10. Balance the following reactions. No fractional coefficients, coefficients must be the lowest whole numbers.

a) 2 K + 2  $\text{H}_2\text{O}$  → 2  $\text{KOH}$  + 1  $\text{H}_2$

b) 2  $\text{H}_2\text{O}_2$  → 2  $\text{H}_2\text{O}$  + 1  $\text{O}_2$

c) 2  $\text{C}_2\text{H}_6$  + 7  $\text{O}_2$  → 4  $\text{CO}_2$  + 6  $\text{H}_2\text{O}$

d) 3  $\text{KOH}$  + 1  $\text{H}_3\text{PO}_4$  → 1  $\text{K}_3\text{PO}_4$  + 3  $\text{H}_2\text{O}$

e) 2  $\text{C}_2\text{H}_5\text{OH}$  + 3  $\text{O}_2$  → 4  $\text{CO}_2$  + 6  $\text{H}_2\text{O}$

11. How many g of  $\text{CO}_2$  are formed from the combustion of 2.9 g of butane,  $\text{C}_4\text{H}_{10}$ , with excess oxygen?

The balanced reaction is:  $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$

- A. 0.26g      B. 6.6g      C. 8.8g      D. 10.5g      E. 2.2g

12. How many g  $\text{I}_2$  will react with 20.4 g Al to form  $\text{AlI}_3$  (aluminum iodide)? As always, first balance the reaction! (Hint: this is not a limiting reagent problem)

- A. 144. g      B. 288. G      C. 192. G      D. 72 g      E. 576 g

13. For the balanced reaction:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ . How many moles of  $\text{NH}_3$  will be produced when 0.50 mol  $\text{N}_2$  reacts with 1.0 mol  $\text{H}_2$ ?

- A. 0.67 mol      B. 1.0 mol      C. 0.50 mol      D. 2.0 mol      E. 3.0 mol

14. Calculate the mass of  $\text{CO}_2$  formed when 2.3 g of ethanol,  $\text{C}_2\text{H}_5\text{OH}$  reacts with 3.2 g oxygen  $\text{O}_2$  to form  $\text{CO}_2$  and  $\text{H}_2\text{O}$ . (You balanced the reaction in 8.e)! This is a limiting reagent problem!

- A. 4.6g      B. 2.2g      C. 4.4g      D. 6.6g      E. 2.93g

# Chem 101 Multiple Choice Questions

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