

Chemistry 100
 Spring 2022

1. A rectangular prism measures 1.2 cm x 0.5 cm x 0.2 cm. It has a mass equal to 4.0 g.

Calculate the volume in cm³ (showed sig figs) $1.2 \times 0.5 \times 0.2 = 0.6 \text{ cm}^3$
 (12) 1.2 cm
 (5) 0.5 cm
 (2) 0.2 cm

Calculate the density in g/cm³

$$\text{Density} = \frac{\text{mass}}{\text{volume}} = \frac{4.0 \text{ g}}{0.6 \text{ cm}^3} = 6.7 \text{ g/cm}^3$$

The density of water is 1.0 g/cm³ and the density of lead is 11.3 g/cm³. Is it more difficult to measure its something else?

$$\frac{1.0 \text{ g/cm}^3}{11.3 \text{ g/cm}^3} = \frac{1}{11.3} \approx 0.088 \text{ (or } 8.8\%)$$

2. A sample is made up of phosphorus and silicon. It has 25.0 g of phosphorus and 12.5 g of silicon. Another sample is made by adding 10.0 g of phosphorus and 20.0 g of silicon and reforming them to form 30.0 g of phosphorus and silicon.

a. Show that these two compounds are different compounds.

$$\frac{25.0 \text{ g P}}{12.5 \text{ g Si}} = 2.0 \quad \frac{30.0 \text{ g P}}{20.0 \text{ g Si}} = 1.5$$

b. Show that these two compounds obey the law of multiple proportions.

$$\frac{2.0}{1.5} = \frac{4}{3} \text{ or } 1.33 \text{ (ratio)}$$

Chem 400 Lab Manual Key

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