Basic Transport Phenomena in Biomedical Engineering 4th Edition Fournier Solutions Manual

Chapter I Solutions

Problem 1.1. Use the conversion factors

$$\frac{8.314J}{\text{mol K}} \times \frac{0.23901\text{cal}}{J} = 1.987 \frac{\text{cal}}{\text{mol K}}$$

Problem 1.2. Use Equation 1.2 to convert the temperature to degrees Celsius

$$t \cdot C = \frac{5}{9} (t \cdot F - 32) = \frac{5}{9} (98.6 - 32) = 37 \cdot C$$

Now, use Equation 1.3 to convert the temperature from degrees Celsius to Kelvin

$$TK = t \cdot C + 273.15 = 37 + 273.15 = 310.15 K$$

Problem 1.3. Use the conversion factors

$$2 \times 10^4 \,\mu dynes \times \frac{dynes}{10^6 \,\mu dynes} \times \frac{N}{105 dynes} \times \frac{kN}{10^3 \,N} = 2 \times 10^{-10} \,kN$$

Problem 1.4. Use the conversion factors

$$\frac{27 \, \mu m}{hr} \times \frac{10^{-6} \, m}{\mu m} \times \frac{0.0006214 miles}{m} \times \frac{hr}{60 min} \times \frac{min}{60 sec} = 4.66 \times 10^{-12} \frac{miles}{sec}$$

Problem 1.5. Use the conversion factors

$$48MW \times \frac{106W}{MW} \times \frac{kg \text{ m}^2}{\text{sec}^3 \text{ W}} \times \frac{10^3 \text{ g}}{\text{kg}} \times \frac{100cg}{g} \times \frac{3600^3 \text{ sec}^3}{\text{hr}^3} \times \frac{100^2 \text{ cm}^2}{\text{m}^2} = 2.24 \times 10^{27} \frac{\text{cgcm}^2}{\text{hr}^3}$$

$$48MW \times \frac{10^6 \text{ W}}{\text{MW}} \times \frac{\text{J}}{\text{sec W}} \times \frac{\text{kJ}}{10^3 \text{J}} \times \frac{60 \text{sec}}{\text{min}} = 2.88 \times 10^6 \frac{\text{kJ}}{\text{min}}$$

Problem 1.6. Use the conversion factors to get both values into units of mm per second and then compare the two values for the higher value.

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Table of Contents Geiger Transport Phenomena Solutions Manual

- 1. Understanding the eBook Geiger Transport Phenomena Solutions Manual
 - The Rise of Digital Reading Geiger Transport Phenomena Solutions Manual
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Geiger Transport Phenomena Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Geiger Transport Phenomena Solutions Manual
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Geiger Transport Phenomena Solutions Manual
 - Personalized Recommendations
 - Geiger Transport Phenomena Solutions Manual User Reviews and Ratings
 - Geiger Transport Phenomena Solutions Manual and Bestseller Lists
- 5. Accessing Geiger Transport Phenomena Solutions Manual Free and Paid eBooks
 - Geiger Transport Phenomena Solutions Manual Public Domain eBooks
 - Geiger Transport Phenomena Solutions Manual eBook Subscription Services
 - Geiger Transport Phenomena Solutions Manual Budget-Friendly Options
- 6. Navigating Geiger Transport Phenomena Solutions Manual eBook Formats
 - o ePub, PDF, MOBI, and More
 - Geiger Transport Phenomena Solutions Manual Compatibility with Devices
 - Geiger Transport Phenomena Solutions Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Geiger Transport Phenomena Solutions Manual
 - Highlighting and Note-Taking Geiger Transport Phenomena Solutions Manual
 - Interactive Elements Geiger Transport Phenomena Solutions Manual
- 8. Staying Engaged with Geiger Transport Phenomena Solutions Manual

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Geiger Transport Phenomena Solutions Manual
- 9. Balancing eBooks and Physical Books Geiger Transport Phenomena Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Geiger Transport Phenomena Solutions Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Geiger Transport Phenomena Solutions Manual
 - Setting Reading Goals Geiger Transport Phenomena Solutions Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Geiger Transport Phenomena Solutions Manual
 - Fact-Checking eBook Content of Geiger Transport Phenomena Solutions Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - o Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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