



Unit 6 Test Study Guide

January 13th, 2022

Brensted-Lowry Acid/Base Theory

- Acids: Proton (H^+) donor
- Bases: Proton (H^+) acceptor
- Conjugate Acids-Base Pair: Two substances that differ only by the presence of an H^+
 - Conjugate Acid: Formed when a proton is added to a base (positive charge)
 - Add a Hydrogen to the formula and add +1 charge
 - Conjugate Base: Formed when a proton is removed from an acid (negative charge)
 - Take away a Hydrogen from the formula and subtract -1 charge
- Amphoteric: The ability for a compound to act as an acid OR as a base (EX: H_2O)
- Acids are classified based on how many H^+ are present
 - Monoprotic: 1 H^+
 - Diprotic: 2 H^+
 - Triprotic: 3 H^+
 - Diprotic and triprotic acids are also amphoteric because they have more than one proton that they can lose

Properties of Strong and Weak Acids/Bases

Strong Acids/Bases	Weak Acids/Bases
<ul style="list-style-type: none"> • Strong electrolytes • Ionize completely in water • Equilibrium is not present <ul style="list-style-type: none"> ◦ Very unlikely that the ions reform into products • Very weak attraction between ions <ul style="list-style-type: none"> ◦ Split apart easily 	<ul style="list-style-type: none"> • Weak electrolytes • Does not ionize completely in water • Equilibrium exists • Strong attraction between ions <ul style="list-style-type: none"> ◦ Do not split apart easily
<p style="text-align: center;">Strong acid</p> <div style="text-align: center;">  </div>	<p style="text-align: center;">Weak acid</p> <div style="text-align: center;">  </div>

Equilibrium And Acid Base Study Guide

David R. Klein



Equilibrium And Acid Base Study Guide:

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Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems With Organic Chemistry Student Study Guide and Solutions Manual 5th Edition students can learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry *Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28

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includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a

glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer *Organic Chemistry, Student Study Guide and Solutions Manual* David R. Klein, 2017-01-04 This is the Student Study Guide and Solutions Manual to accompany Organic Chemistry 3e Organic Chemistry 3rd Edition is not merely a compilation of principles but rather it is a disciplined method of thought and analysis Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems Readers must learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry Existing textbooks provide extensive coverage of the principles but there is far less emphasis on the skills needed to actually solve problems [Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry](#) Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section's critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

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