



Chemical Systems in Equilibrium

In this chapter, you will be able to

- demonstrate an understanding of dynamic equilibrium in a variety of situations, and use the law of chemical equilibrium to solve problems;
- apply Le Châtelier's Principle to predict how various factors affect a chemical system at equilibrium;
- carry out experiments to determine equilibrium constants;
- explain how equilibrium principles may be applied to optimize the production of industrial chemicals;
- describe the behaviour of ionic solutes in solutions that are unsaturated, saturated, and supersaturated;
- calculate the molar solubility of a pure substance in water or in a solution of a common ion, and predict the formation of precipitates by using the solubility product constant;
- identify the entropy changes associated with chemical and physical processes and the tendency of reactions to achieve minimum energy and maximum entropy;
- identify the effects of solubility on biological systems.

Have you ever felt as though you were at rest while walking up the “down” escalator, or as you kept pace with a speeding treadmill? If so, you have experienced a state of *dynamic equilibrium*—a situation where at least one property remains constant while opposing processes occur at the same rate. On the escalator or the treadmill, your body's position is the property that remains constant, while your legs and the platform move in opposite directions at the same speed.

Consider a swimming pool. Water jets on the sides of the pool continuously inject warm, clean water into the pool, but the water level remains constant. For this to occur, water must be leaving the pool through drains at the same rate as it is entering through the jets. The water in the pool is in a state of dynamic equilibrium. Water molecules are entering and leaving the pool at the same rate while the water level remains constant.

There are many examples of dynamic equilibrium in the world of chemistry. Although they are not exactly the same as the situations described above, there are many similarities. Chemical systems at equilibrium all display at least one constant property. For example, the colour of a reaction mixture could remain constant, giving the impression that nothing is happening. This can be deceiving. In the pool example, close inspection reveals competing activities occurring at the same rate: water entering and water leaving. What competing activities could be occurring in a chemical reaction at equilibrium?



REFLECT on your learning

- Two children are perfectly balanced on a teeter-totter. Is this a case of dynamic equilibrium? Explain.
- Describe two household examples of dynamic equilibrium.
- The molar solubility of $\text{PbI}_{2(s)}$ in pure water is $1.3 \times 10^{-3} \text{ mol/L}$. In water that contains $0.1 \text{ mol/L NaI}_{(aq)}$, the solubility of $\text{PbI}_{2(s)}$ is $7.9 \times 10^{-7} \text{ mol/L}$. Provide reasons for the difference in the two solubilities.
- (a) What does the double arrow in the following equation mean?

$$2 \text{CO}_{(g)} + \text{O}_{2(g)} \rightleftharpoons 2 \text{CO}_{2(g)} \quad \Delta H^\circ = -566.1 \text{ kJ}$$
 (b) If the above reaction occurs in a closed container, would heating the container increase or decrease the production of $\text{CO}_{2(g)}$? Explain.
- Will the following reaction have a tendency to occur spontaneously as written (at SATP)? Explain.



Characteristics Of Chemical Equilibrium Lab Answers

Lei Huang



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Table of Contents Characteristics Of Chemical Equilibrium Lab Answers

1. Understanding the eBook Characteristics Of Chemical Equilibrium Lab Answers
 - The Rise of Digital Reading Characteristics Of Chemical Equilibrium Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Characteristics Of Chemical Equilibrium Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Characteristics Of Chemical Equilibrium Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Characteristics Of Chemical Equilibrium Lab Answers
 - Personalized Recommendations
 - Characteristics Of Chemical Equilibrium Lab Answers User Reviews and Ratings
 - Characteristics Of Chemical Equilibrium Lab Answers and Bestseller Lists
5. Accessing Characteristics Of Chemical Equilibrium Lab Answers Free and Paid eBooks
 - Characteristics Of Chemical Equilibrium Lab Answers Public Domain eBooks
 - Characteristics Of Chemical Equilibrium Lab Answers eBook Subscription Services
 - Characteristics Of Chemical Equilibrium Lab Answers Budget-Friendly Options
6. Navigating Characteristics Of Chemical Equilibrium Lab Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Characteristics Of Chemical Equilibrium Lab Answers Compatibility with Devices
 - Characteristics Of Chemical Equilibrium Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Characteristics Of Chemical Equilibrium Lab Answers
 - Highlighting and Note-Taking Characteristics Of Chemical Equilibrium Lab Answers
 - Interactive Elements Characteristics Of Chemical Equilibrium Lab Answers
 8. Staying Engaged with Characteristics Of Chemical Equilibrium Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Characteristics Of Chemical Equilibrium Lab Answers
 9. Balancing eBooks and Physical Books Characteristics Of Chemical Equilibrium Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Characteristics Of Chemical Equilibrium Lab Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Characteristics Of Chemical Equilibrium Lab Answers
 - Setting Reading Goals Characteristics Of Chemical Equilibrium Lab Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Characteristics Of Chemical Equilibrium Lab Answers
 - Fact-Checking eBook Content of Characteristics Of Chemical Equilibrium Lab Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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