

Name _____ Date _____ Class _____

CHAPTER 4 STUDY GUIDE FOR CONTENT MASTERY**The Structure of the Atom****Section 4.1 Early Theories of Matter***In your textbook, read about the philosophers, John Dalton, and defining the atom.**For each statement below, write true or false.*

- false** _____ 1. Ancient philosophers regularly performed controlled experiments.
- true** _____ 2. Philosophers formulated explanations about the nature of matter based on their own experiences.
- true** _____ 3. Both Democritus and Dalton suggested that matter is made up of atoms.
- true** _____ 4. Dalton's atomic theory stated that atoms separate, combine, or rearrange in chemical reactions.
- false** _____ 5. Dalton's atomic theory stated that matter is mostly empty space.
- false** _____ 6. Dalton was correct in thinking that atoms could not be divided into smaller particles.
- true** _____ 7. Dalton's atomic theory stated that atoms of different elements combine in simple whole-number ratios to form compounds.
- true** _____ 8. Dalton thought that all atoms of a specific element have the same mass.
- false** _____ 9. Democritus proposed that atoms are held together by chemical bonds, but no one believed him.
- true** _____ 10. Dalton's atomic theory was based on careful measurements and extensive research.
- false** _____ 11. There are no instruments powerful enough to measure atoms so that they can be seen.
- true** _____ 12. The smallest particle of an element that retains the properties of that element is called an atom.

Study Guide for Content Mastery

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CHAPTER 4 STUDY GUIDE FOR CONTENT MASTERY**Section 4.2 Subatomic Particles and the Nuclear Atom***In your textbook, read about discovering the electron and the nuclear atom.**For each item in Column A, write the letter of the matching item in Column B.*

| Column A | Column B |
|--|---------------|
| 5. _____ 1. Proposed the nuclear atomic model | a. Thomson |
| 6. _____ 2. Determined the mass-to-charge ratio of an electron | b. Millikan |
| 7. _____ 3. Calculated the mass of an electron | c. Rutherford |

Draw and label a diagram of each atomic model.

4. **plum pudding model**

Drawing should look like a ball of chocolate chip cookie dough. The chocolate chips should be labeled with negative charge or as electrons. The dough should be labeled as evenly distributed positive charges.

5. **no clear atomic model**

Drawing should look like a peach with a pit. The pit should be labeled nucleus and should include labeled protons and neutrons. The outer circle of the peach should be labeled electrons.

*In your textbook, read about the discovery of protons and neutrons.**Complete the following table of protons, electrons, and neutrons characteristics.*

| Particle | Symbol | Location | Relative Charge | Relative Mass |
|-------------|--------|--------------------------------------|-----------------|---------------|
| 1. Proton | p^+ | In the nucleus | $1+$ | 1 |
| 2. Neutron | n^0 | In the nucleus | 0 | 1 |
| 3. Electron | e^- | In the space surrounding the nucleus | $1-$ | $1/1840$ |

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Study Guide for Content Mastery

Chapter 4 Study Guide Answer Key Physics

Arshad Iqbal



Chapter 4 Study Guide Answer Key Physics:

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