

Chapter 15: Acids and Bases



Acids and Bases

Arrhenius Definitions:

- acids - compounds that produce an increase in $[H^+]$ when dissolved in water
- bases - compounds that produce an increase in $[OH^-]$ when dissolved in water

Lewis Definitions:

- acids - electron pair acceptors
- bases - electron pair donors

Bronsted-Lowry Definitions:

- acids - H^+ donors
- bases - H^+ acceptors

Lewis Acids & Lewis Bases

- more broad way to define acids and bases
- Lewis acids – electron pair acceptors
metal cations (M^{n+}) and boron are common Lewis acids
species that are electron deficient; electrophiles
- Lewis bases – electron pair donors
species with O, N, halogen frequently have lone pairs of electrons to share \therefore Lewis bases
species that are electron rich; nucleophiles
- product of a Lewis Acid + Lewis Base reaction is called a Lewis Acid-Base adduct

Lewis Acids & Lewis Bases

- some examples:



- acidic oxides (oxides of nonmetals):



- basic oxides (oxides of metals):



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