

The Mole

Section 10.1 Measuring Matter

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Practice Problems

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1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

$$2.50 \text{ mol-Zn} \times \frac{6.02 \times 10^{23} \text{ atoms Zn}}{1 \text{ mol-Zn}} \\ = 1.51 \times 10^{24} \text{ atoms of Zn}$$

2. Calculate the number of molecules in 11.5 mol of water (H₂O).

$$11.5 \text{ mol-H}_2\text{O} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol-H}_2\text{O}} \\ = 6.92 \times 10^{24} \text{ molecules of H}_2\text{O}$$

3. Silver nitrate (AgNO₃) is used to make several different silver halides used in photographic films. How many formula units of AgNO₃ are there in 3.25 mol AgNO₃?

$$3.25 \text{ mol-AgNO}_3 \times \frac{6.02 \times 10^{23} \text{ formula units AgNO}_3}{1 \text{ mol-AgNO}_3} \\ = 1.96 \times 10^{24} \text{ formula units of AgNO}_3$$

4. **Challenge** Calculate the number of oxygen atoms in 5.0 mol of oxygen molecules. Oxygen is a diatomic molecule, O₂.

$$5.00 \text{ mol-O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules-O}_2}{1 \text{ mol-O}_2} \\ \times \frac{2 \text{ O atoms}}{1 \text{ molecule-O}_2} = 6.02 \times 10^{24} \text{ atoms O}$$

5. How many moles contain each of the following?

a. 5.75×10^{24} atoms Al

$$5.75 \times 10^{24} \text{ atoms-Al} \times \frac{1 \text{ mol Al}}{6.02 \times 10^{23} \text{ atoms-Al}} \\ = 9.55 \text{ mol Al}$$

b. 2.50×10^{20} atoms Fe

$$2.50 \times 10^{20} \text{ atoms-Fe} \times \frac{1 \text{ mol Fe}}{6.02 \times 10^{23} \text{ atoms-Fe}} \\ = 4.15 \times 10^{-4} \text{ mol Fe}$$

6. **Challenge** Identify the representative particle for each formula and convert the given number of representative particles to moles.

a. 3.75×10^{24} CO₂

The representative particle is a molecule.

$$3.75 \times 10^{24} \text{ molecules-CO}_2 \\ \times \frac{1 \text{ mol CO}_2}{6.02 \times 10^{23} \text{ molecules-CO}_2} = 6.23 \text{ mol CO}_2$$

b. 3.58×10^{23} ZnCl₂

The representative particle is a formula unit.

$$3.58 \times 10^{23} \text{ formula-units-ZnCl}_2 \\ \times \frac{1 \text{ mol ZnCl}_2}{6.02 \times 10^{23} \text{ formula-units-ZnCl}_2} \\ = 0.595 \text{ mol ZnCl}_2$$

Section 10.1 Assessment

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7. **Explain** why chemists use the mole.

Chemists use the mole because it is a convenient way of knowing how many representative particles are in a sample.

8. **State** the mathematical relationship between Avogadro's number and 1 mol.

One mole contains 6.02×10^{23} representative particles.

9. **List** the conversion factors used to convert between particles and moles.

$$\frac{6.02 \times 10^{23} \text{ representative particles}}{1 \text{ mol}} \\ \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ representative particles}}$$

Chapter 10 Study Guide The Mole Answer Key

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not administer these meetings, please use the listing contacts for any questions. 12 Steps of LAA (Love Addicts Anonymous) - 12Step.org Sought through prayer and meditation to improve our conscious contact with God as we understood God, praying only for knowledge of God's will for us and the ... Sex and Love Addicts Anonymous (S.L.A.A.) - Fellowship ... The S.L.A.A. F.W.S. BOT encourages all S.L.A.A. members to value our differences and bring our authentic, whole selves to the rooms. Our diverse voices bring ... Sex and Love Addicts Anonymous Sex and Love Addicts Anonymous (SLAA) is a twelve-step program for people recovering from sex addiction and love addiction. SLAA was founded in Boston, ... LAA stands for Love Addicts Anonymous This definition appears very frequently and is found in the following Acronym Finder categories: Organizations, NGOs, schools, universities, etc. LAA Step Guide by Love Addicts Anonymous : \$15.94 May 17, 2023 — This Twelve Steps Guide is the result of the long-term work of our group consciousness and our experience in working the Steps. Love Addicts Anonymous Love Addicts Anonymous, San Francisco Bay Area. 757 likes · 5 talking about this. Love Addicts Anonymous is a twelve step program for love addicts.