

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 15** **STUDY GUIDE FOR CONTENT MASTERY**

## Solutions

### Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is air! (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen is air is the (2) solvent and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is air is (4) soluble substances. A substance that does not dissolve is (5) insoluble. (6) liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



8. Hydration is solvation in which the solvent is water. What is solvation?

**Solvation is the process of surrounding solute particles with solvent particles to form a solution.**

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85

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 15** **STUDY GUIDE FOR CONTENT MASTERY**

### Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?

**The sodium and chloride ions are separated and surrounded by the water molecules.**

11. Explain the attraction of the water molecules around the sodium ions and chloride ions.

**Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.**

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

**The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.**

13. List three ways that the rate of solvation may be increased.

**stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent**

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

**true** 14. The overall energy change that occurs when a solution forms is called the heat of solution.

**false** 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

**false** 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

**true** 17. In a saturated solution, solvation and crystallization are in equilibrium.

**true** 18. Additional solute can be dissolved in an unsaturated solution.

**true** 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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86

# Chapter 26 Study Guide For Content Mastery Answer Key

**Doni L. Bird, Debbie S. Robinson**



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### **Table of Contents Chapter 26 Study Guide For Content Mastery Answer Key**

1. Understanding the eBook Chapter 26 Study Guide For Content Mastery Answer Key
  - The Rise of Digital Reading Chapter 26 Study Guide For Content Mastery Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 26 Study Guide For Content Mastery Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 26 Study Guide For Content Mastery Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 26 Study Guide For Content Mastery Answer Key

- Personalized Recommendations
  - Chapter 26 Study Guide For Content Mastery Answer Key User Reviews and Ratings
  - Chapter 26 Study Guide For Content Mastery Answer Key and Bestseller Lists
5. Accessing Chapter 26 Study Guide For Content Mastery Answer Key Free and Paid eBooks
- Chapter 26 Study Guide For Content Mastery Answer Key Public Domain eBooks
  - Chapter 26 Study Guide For Content Mastery Answer Key eBook Subscription Services
  - Chapter 26 Study Guide For Content Mastery Answer Key Budget-Friendly Options
6. Navigating Chapter 26 Study Guide For Content Mastery Answer Key eBook Formats
- ePub, PDF, MOBI, and More
  - Chapter 26 Study Guide For Content Mastery Answer Key Compatibility with Devices
  - Chapter 26 Study Guide For Content Mastery Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chapter 26 Study Guide For Content Mastery Answer Key
  - Highlighting and Note-Taking Chapter 26 Study Guide For Content Mastery Answer Key
  - Interactive Elements Chapter 26 Study Guide For Content Mastery Answer Key
8. Staying Engaged with Chapter 26 Study Guide For Content Mastery Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 26 Study Guide For Content Mastery Answer Key
9. Balancing eBooks and Physical Books Chapter 26 Study Guide For Content Mastery Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 26 Study Guide For Content Mastery Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 26 Study Guide For Content Mastery Answer Key
- Setting Reading Goals Chapter 26 Study Guide For Content Mastery Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 26 Study Guide For Content Mastery Answer Key

- Fact-Checking eBook Content of Chapter 26 Study Guide For Content Mastery Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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