ANSWERS

Nuclear Chemistry Review Questions

1. Compare the mass, charge, and penetration power of alpha, beta and gamma radiation.

	Mass	Charge	Penetration
Alpha	4	+	Lowest
Beta	0	-	Middle
Gamma	0	0	Highest

2. In 5.49 seconds, 1.20g of argon-35 decays to leave only 0.15g. What is the half-life?

   1.20g → 0.60g → 0.30g → 0.15g , therefore it took 3 half-life periods.

   5.49÷3 = 1.83 sec

3. How many days does it take 16.0g of gold-198 to decay to 1.0g?

   16.0g → 8. 0g → 4.0g → 2.0g → 1.0g , therefore 4 half-life periods

   From Ref Table N, 1 half-life = 2.69d
   2.69d * 4 = 10.76 days

4. Why does an atom undergo radioactive decay?

   Because it has an unstable nucleus

5. Compare and contrast fission and fusion. Why is fission used for electrical and fusion is not?

   Fission=Division, the splitting of large unstable atoms(nucleus)
   Fusion=joining together, small isotopes of H, creates way more energy

6. How is nuclear chemistry applied to medicine? What type of radioisotope must be used when it is injected into a person?

   Radioisotopes are used to treat and diagnose illnesses
   If injected they must have short half-lives and quickly removed from the body

7. Write out the nuclear decay of radium-226.

   $^{226}_{88}Ra \rightarrow {}^{4}_{2}He + {}^{222}_{86}Rn$

8. What is different about nuclear reactions when compared to a chemical reaction?

   Nuclear reactions transform atoms into other atoms, chemical reactions involved the rearrangement of electrons

9. At what point do all isotopes of an element become radioactive?

   Up to and including atomic # 83

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SJ Ball

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