

10**CHEMICAL QUANTITIES****Practice Problems**

On a piece of paper, solve the following problems.

YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER

1. What is the molar mass of sucrose ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$)?
2. What is the molar mass of each of the following compounds?
 - a. phosphorus pentachloride (PCl_5)
 - b. uranium hexafluoride (UF_6)
3. Calculate the molar mass of each of the following ionic compounds:
 - a. KMnO_4
 - b. $\text{Ca}_3(\text{PO}_4)_2$
4. How many moles is 3.52×10^{24} molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen (O_2)?

SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS

1. What is the molar mass of each of the following compounds?
 - a. $\text{C}_6\text{H}_{12}\text{O}_6$
 - b. NaHCO_3
 - c. C_3H_{12}
 - d. KNH_4SO_4
2. Calculate the mass in grams of each of the following:
 - a. 8.0 mol lead oxide (PbO)
 - b. 0.75 mol hydrogen sulfide (H_2S)
 - c. 0.00100 mol silicon tetrahydride (SiH_4)
 - d. 1.50×10^{-2} mol molecular oxygen (O_2)
 - e. 2.30 mol ethylene glycol ($\text{C}_2\text{H}_6\text{O}_2$)
3. How many grams are in 1.73 mol of dinitrogen pentoxide (N_2O_5)?
4. How many grams are in 0.658 mol of calcium phosphate [$\text{Ca}_3(\text{PO}_4)_2$]?
5. Calculate the number of moles in each of the following:
 - a. 0.50 g sodium bromide (NaBr)
 - b. 13.5 g magnesium nitrate [$\text{Mg}(\text{NO}_3)_2$]
 - c. 1.02 g magnesium chloride (MgCl_2)
 - d. 0.00100 g monochloromethane (CH_3Cl)
 - e. 1.50×10^{-3} g propylene glycol [$\text{C}_3\text{H}_8(\text{OH})_2$]
6. A chemist plans to use 435.0 grams of ammonium nitrate (NH_4NO_3) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ($\text{C}_{20}\text{H}_{24}\text{N}_2\text{O}_2$). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane (CH_4) gas?
9. How many moles is 135 L of ammonia (NH_3) gas at STP?

Chapter 10 Chemical Quantities Study Guide Answer Key

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Table of Contents Chapter 10 Chemical Quantities Study Guide Answer Key

1. Understanding the eBook Chapter 10 Chemical Quantities Study Guide Answer Key
 - The Rise of Digital Reading Chapter 10 Chemical Quantities Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 10 Chemical Quantities Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 10 Chemical Quantities Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 10 Chemical Quantities Study Guide Answer Key
 - Personalized Recommendations
 - Chapter 10 Chemical Quantities Study Guide Answer Key User Reviews and Ratings
 - Chapter 10 Chemical Quantities Study Guide Answer Key and Bestseller Lists
5. Accessing Chapter 10 Chemical Quantities Study Guide Answer Key Free and Paid eBooks
 - Chapter 10 Chemical Quantities Study Guide Answer Key Public Domain eBooks
 - Chapter 10 Chemical Quantities Study Guide Answer Key eBook Subscription Services
 - Chapter 10 Chemical Quantities Study Guide Answer Key Budget-Friendly Options

6. Navigating Chapter 10 Chemical Quantities Study Guide Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 10 Chemical Quantities Study Guide Answer Key Compatibility with Devices
 - Chapter 10 Chemical Quantities Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 10 Chemical Quantities Study Guide Answer Key
 - Highlighting and Note-Taking Chapter 10 Chemical Quantities Study Guide Answer Key
 - Interactive Elements Chapter 10 Chemical Quantities Study Guide Answer Key
8. Staying Engaged with Chapter 10 Chemical Quantities Study Guide Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 10 Chemical Quantities Study Guide Answer Key
9. Balancing eBooks and Physical Books Chapter 10 Chemical Quantities Study Guide Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 10 Chemical Quantities Study Guide Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 10 Chemical Quantities Study Guide Answer Key
 - Setting Reading Goals Chapter 10 Chemical Quantities Study Guide Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 10 Chemical Quantities Study Guide Answer Key
 - Fact-Checking eBook Content of Chapter 10 Chemical Quantities Study Guide Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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