

4. A sample of oxygen gas has a volume of 150 mL when its pressure is 440 mmHg. If the pressure is increased to standard pressure and the temperature remains constant, what will the new gas volume be?

Boyle's Law

$$P_1 V_1 = P_2 V_2$$

$$V_1 = 150 \text{ mL}$$

$$P_1 = 440 \text{ mmHg}$$

$$V_2 = ?$$

$$P_2 = 760 \text{ mmHg (1 atm)}$$

$$(440 \text{ mmHg})(150 \text{ mL}) = V_2 (760 \text{ mmHg})$$

$$V_2 = 87 \text{ mL}$$

5. Ralph had a helium balloon with a volume of 4.88 liters at 150 kPa of pressure. If the volume is changed to 3.15 liters, what would be the new pressure in atm?

Boyle's Law

$$P_1 V_1 = P_2 V_2$$

$$P_1 = 150 \text{ kPa}$$

$$V_1 = 4.88 \text{ L}$$

$$P_2 = ?$$

$$V_2 = 3.15 \text{ L}$$

$$(4.88 \text{ L})(150 \text{ kPa}) = (3.15 \text{ L}) P_2$$

$$P_2 = 232 \text{ kPa} \times \frac{1 \text{ atm}}{101.3 \text{ kPa}}$$

$$P_2 = 2.29 \text{ atm}$$

6. 5.36 liters of nitrogen gas are at -25°C and 733 mm Hg. What would be the volume at 128°C and 1.5 atm?

Combined gas Law

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$P_1 = 733 \text{ mmHg} \times \frac{1 \text{ atm}}{760 \text{ mmHg}} = 0.964 \text{ atm}$$

$$V_1 = 5.36 \text{ L}$$

$$P_2 = 1.5 \text{ atm}$$

$$T_1 = -25^\circ\text{C} + 273 = 248 \text{ K}$$

$$T_2 = 128^\circ\text{C} + 273 = 401 \text{ K}$$

$$\frac{(0.964 \text{ atm})(5.36 \text{ L})}{248 \text{ K}} = \frac{V_2 (1.5 \text{ atm})}{401 \text{ K}}$$

$$V_2 = 5.57 \text{ L}$$

7. At constant temperature, 245 mL of a gas at 4 atm of pressure is expanded to 1.75 L. What is the new pressure?

Boyle's Law

$$P_1 V_1 = P_2 V_2$$

$$V_1 = 245 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 0.245 \text{ L}$$

$$P_1 = 4 \text{ atm}$$

$$V_2 = 1.75 \text{ L}$$

$$(0.245 \text{ L})(4 \text{ atm}) = (1.75 \text{ L}) P_2$$

$$P_2 = 0.56 \text{ atm}$$

Charles And Boyles Law Gizmo Answer Key

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