9.1

Objectives

- Define Stoichiometry
- Describe the importance of the mole ratio in stoichiometric calculations
- Write a mole ratio relating two substances in a chemical reaction.

In stoichiometry problems you will be given the amount of at least one of the substances in a chemical equation and asked to find the amount of one or more substances in the equation

The given amount may be expressed in ______ or _____

The Unknown amount may be expressed in ______ or ______ or ______

What are the 4 types of stoichiometry problems?

What conversion factor do you use to change grams to moles?

All stoichiometry problems involve the use of the

Where do you get the mole ratio from?

What does the mole ratio enable you to do?

Can you perform a stoichiometry problem with an unbalanced chemical equation? Explain

9.1 assignment p 285 1 to 4 p 305 1 to 3 ws pages 1.2

9.2

Objectives

- Calculate the amount in <u>moles</u> of a reactant or product from the amount in <u>moles</u> of a different reactant or product
- Calculate the mass of a reactant or product from the amount in moles of a different reactant or product
- Calculate the amount in moles of a reactant or product from the mass of a different reactant or product
- Calculate the mass of a reactant or product from the mass of a different mass or product

Given moles, asked to find moles, use mole ratio

Do p 289, 1,2 Here are the balanced equations 3H2 + N2 →2NH3, 2KClO3 → 2KCl + 3O2

Given mole, asked to find grams, Use molar mass to convert grams to moles, then use mole Ratio

Do p 291, 1 and 2

Given grams, asked to find moles, use molar mass to convert grams to moles, then mole ratio

Do p 293, 1,2 Here is the balanced equation 2HgO → 2Hg + O2

<u>Chapter 9 Stoichiometry Guided Reading And Study</u> <u>Workbook Answers</u>

David Baud

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